

Study Guide Chapter 9 – Covalent Bonding

1. Describe covalent bonds and be able to distinguish between single, double, triple, sigma and pi bonds.
2. What determines the strength of covalent bonds?
3. Be able to draw Lewis structures (or electron dot diagrams) for Mg, Li, C, Al, S, Br₂, MgO, HF, CH₄, Li₂S, and SO₃. Name the structures and label ionic or covalent.
4. Be able to write and name molecular compounds and acids.
5. Draw the difference between simple formulas, Lewis structures, and structural formulas?
6. Give examples of exceptions to the octet rule including electron deficient molecules, expanded octets, and coordinated covalent bonds.
7. Describe hybridization and give the energy level diagram for carbon.
8. Be able to identify the seven different VSEPR models (or geometry). Build and draw the geometry of H₂S, CO, AlF₃, PCl₃, MgBr₂, PCl₅ and SBr₆. State how many bonded pairs, lone pairs, and bond angles for each.
9. Draw SO₃ and describe why it is a resonance structure.
10. Can you discuss and draw the formation of an ester?
11. Draw methyl alcohol and formic acid then draw the ester and name it.
12. What are esters used for?
13. Draw a diagram of nonpolar covalent, polar covalent and ionic bonds. Distinguish between them with one sentence.
14. Be able to use electronegativities to distinguish between nonpolar, polar covalent and ionic bonds.
15. Describe how you would identify a molecules polarity.
16. Describe some properties of covalent compounds and solids.

Chapter 9 Covalent Bonding Study Guide Answers

Patrick Vollmar



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Unveiling the Magic of Words: A Report on "**Chapter 9 Covalent Bonding Study Guide Answers**"

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