

Name _____ Date _____ Class _____

CHAPTER 9 STUDY GUIDE FOR CONTENT MASTERY

Covalent Bonding

Section 9.1 The Covalent Bond

In your textbook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the passage.

covalent bond	molecule	sigma bond	exothermic	pi bond
---------------	----------	------------	------------	---------

When sharing of electrons occurs, the attachment between atoms that results is called a(n) **(1) covalent bond**. When such an attachment is formed, bond dissociation energy is released, and the process is **(2) exothermic**. When two or more atoms bond by means of electron sharing, the resulting particle is called a(n) **(3) molecule**. If the electrons shared are contained between the two atoms, the attachment is called a(n) **(4) sigma bond**. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n) **(5) pi bond**.

In your textbook, read about single and multiple bonds and bond strengths.

Circle the letter of the choice that best completes the statement or answers the question.

- Is what form do elements such as hydrogen, nitrogen, and oxygen normally occur?
 - a single atom
 - a molecule containing three atoms**
 - a molecule containing two atoms
 - a molecule containing four atoms
- How many electrons are shared in a double covalent bond?
 - none
 - one
 - two
 - four**
- Bond lengths in the distance between
 - two molecules of the same substance
 - the nuclei of two attached atoms**
 - the electrons in two attached atoms
 - the orbitals of two attached atoms
- Which of the following relationships relating to bond length is generally correct?
 - the shorter the bond, the stronger the bond**
 - the shorter the bond, the weaker the bond
 - the shorter the bond, the fewer the electrons in it
 - the shorter the bond, the lower the bond dissociation energy

Study Guide for Content Mastery

Chemistry: Matter and Change • Chapter 9

49

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CHAPTER 9 STUDY GUIDE FOR CONTENT MASTERY

Section 9.2 Naming Molecules

In your textbook, read about how binary compounds and acids are named from their formulas.

For each statement below, write *true* or *false*.

- Binary molecular compounds are generally composed of a metal and a nonmetal.
false
- The second element in the formula of a binary compound is named using the suffix *-ite*.
false
- The prefix *tri-* indicates three atoms.
true
- The prefix *kilo-* indicates six atoms.
true
- In naming the first element in a formula, the prefix *mono-* is not used.
true
- For binary acids, the hydrogen part of the compound is named using the prefix *hydro-*.
false
- An oxyacid contains only two elements.
false
- If the name of the cation of an oxyacid ends in *-ite*, the acid name contains the suffix *-ous*.
true

In your textbook, read about naming molecular compounds and oxyacids.

For each item in Column A, write the letter of the matching item in Column B.

Column A	Column B
5. Cl_2	a. hydrobromic acid
6. CO	b. nitrogen trioxide
7. H_2O	c. carbon monoxide
8. H_2CO_3	d. nitrous acid
9. NO_2	e. acetone
10. N_2O_5	f. nitric acid
11. HNO_3	g. carbonic acid
12. HNO_2	h. boric acid
13. HBr	i. carbon dioxide
14. HBrO_3	

50 Chemistry: Matter and Change • Chapter 9 Study Guide for Content Mastery

Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons

TD Snyder



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Table of Contents Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons

1. Understanding the eBook Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons
 - The Rise of Digital Reading Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons
 - Personalized Recommendations
 - Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons User Reviews and Ratings
 - Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons and Bestseller Lists
5. Accessing Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons Free and Paid eBooks
 - Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons Public Domain eBooks
 - Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons eBook Subscription Services
 - Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons Budget-Friendly Options

6. Navigating Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons Compatibility with Devices
 - Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons
 - Highlighting and Note-Taking Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons
 - Interactive Elements Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons
8. Staying Engaged with Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons
9. Balancing eBooks and Physical Books Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons
 - Setting Reading Goals Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons
 - Fact-Checking eBook Content of Chapter 22 Study Guide For Content Mastery Answer Key Hydrocarbons
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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