Name ______ Date _____ Period ____

Chapter 17 Thermochemistry Study Guide

17.1 - 17.2 Thermochemical Equations

Make the following conversions:

- a. 444 cal to joules = 1.86 x 10³ J
- b. 1.8 kJ to joules = 1.8 x 103 J
- c. 0.45 kJ to calories = 1.1 x 10² cal
- Classify each of these processes as endothermic or exothermic:
 - a. condensing steam exo

c. evaporating alcohol – endo

b. burning alcohol – exo

- d. baking a potato endo
- The specific heat capacity for silver is 0.24 J/g°C. Calculate the energy required to raise the temperature of 150.0 g Ag from 273 K to 298 K. Calculate the molar heat capacity of silver.

Energy: q = (150.0)(.24)(298-273) = 9.0 x 10² J

Molar heat capacity: (0.24 J/g°C)(107.87 g/mol Ag) = 26 J/mol

 It takes 585 J of energy to raise the temperature of 125.6 g Hg from 20.0°C to 53.5°C. Calculate the specific heat capacity and the molar heat capacity of Hg.

Specific heat capacity: C = q/mΔt = 585/(125.6)(53.5-20) = 0.139 J/g°C

Molar heat capacity: $(0.139 \text{ J/g}^{\circ}\text{C})(200.59 \text{ g/mol Hg}) = 27.9 \text{ J/mol}$

 A 46.2-g sample of copper is heated to 95.4°C and then placed in a calorimeter containing 75.0 g water at 19.6°C. The equilibrium temperature in the calorimeter is 21.8°C. Calculate the specific heat capacity of copper; assuming that all the heat lost by the copper is gained by the water.

C = 0.203 J/g°C

 A 15.0-g sample of nickel metal is heated to 100.0°C and dropped into 55.0 g of water, initially at 23.0°C. Assuming that all the heat lost by the nickel is absorbed by the water; calculate the final temperature of the nickel and the water. The specific heat of nickel is 0.444 J/g°C.

> $15.0(.444)(100.0 - T_f)$ = $55.0(4.184)(T_f - 23.0)$ $666 - 6.66T_f$ = $230.12T_f - 5292.76$ 5958.76 = $236.78 T_f$ $25.2^{\circ}C$ = T_f

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JE Gale

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