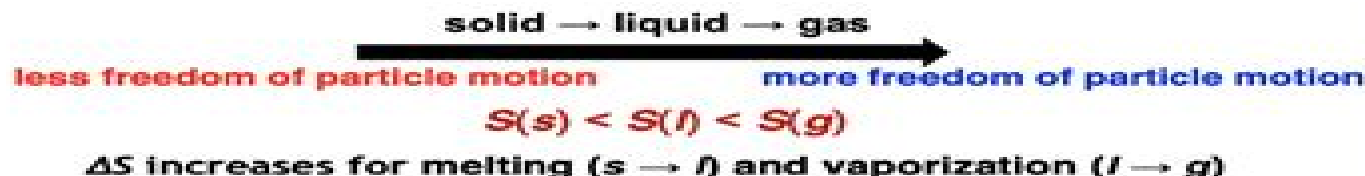


## General Chemistry II Chapter 15 Study Guide

### Entropy:

- Entropy ( $S$ ): is related to freedom of motion of the particles in the system



**Dissolving of soluble salts: solid + liquid  $\rightarrow$  ions in solution**

**localized energy of motion  $\rightarrow$  dispersed energy of motion**

All spontaneous endothermic processes result in an increase  
in *entropy* in the system.

### Microstates and entropy:

- The entropy ( $S$ ) of the system is also defined as the number of microstates ( $W$ ) in a system
- $S = k \ln W$ 
  - $k$  = Boltzmann constant ( $1.38 \times 10^{-23}$  J/K)
  - $W$  = # of microstates
- A system with fewer microstates has lower entropy; a system with more microstates has higher entropy
- Entropy, like enthalpy, is a state function and only depends on the initial and final states
  - $\Delta S = S_f - S_i$ 
    - $S_f$  = final entropy
    - $S_i$  = initial entropy

### Spontaneous expansion of gas:

- When the stopcock is opened, the gas spontaneously expands to fill both flasks
- Increasing the volume increases the number of translational energy levels the particles can occupy; the number of microstates—and the entropy—increases

### Factors affecting Entropy ( $S$ ):

- Entropy of a substance is higher when:
  - Temperature of a substance is increased
  - # of substance particles is increased

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