

CHAPTER 14 PERIODIC PATTERNS IN THE MAIN-GROUP ELEMENTS

END-OF-CHAPTER PROBLEMS

- 14.1** Ionization energy is defined as the energy required to remove the outermost electron from an atom. The further the outermost electron is from the nucleus, the less energy is required to remove it from the attractive force of the nucleus. In hydrogen, the outermost electron is in the $n = 1$ level and in lithium the outermost electron is in the $n = 2$ level. Therefore, the outermost electron in lithium requires less energy to remove, resulting in a lower ionization energy.

- 14.2 ☐ I = ionic M = metallic C = covalent

[illegible]

- 14.3 Plan: Recall that to form hydrogen bonds a compound must have H directly bonded to either N, O, or F

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a) NH_3 will hydrogen bond because H is bonded to N



b) $\text{CH}_3\text{CH}_2\text{OH}$ will hydrogen bond since H is bonded to O. CH_3OCH_3 has no O-H bonds, only C-H bonds.



Chapter 14 Solutions Manual Chemistry

Jason Powell



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