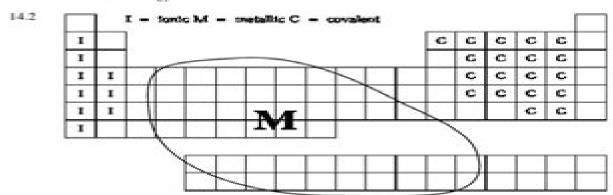
CHAPTER 14 PERIODIC PATTERNS IN THE MAIN-GROUP ELEMENTS

END-OF-CHAPTER PROBLEMS

Ionization energy is defined as the energy required to remove the outermost electron from an atom. The further 14.1 the outermost electron is from the nucleus, the less energy is required to remove it from the attractive force of the nucleus. In hydrogen, the outermost electron is in the n = 1 level and in lithium the outermost electron is in the m = 2 level. Therefore, the outermost electron in lithium requires less energy to remove, resulting in a lower ionization energy.



- 14.3 Plan: Recall that to form hydrogen bonds a compound must have H directly bonded to either N, O, or F. Solution:

 a) NH₃ will hydrogen bond because H is bonded to N.

b) CH₃CH₂OH will hydrogen bond since H is bonded to O. CH₃OCH₃ has no O-H bonds, only C-H bonds.

Chapter 14 Solutions Manual Chemistry

Jason Powell

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