

TEACHER GUIDE AND ANSWERS

Study Guide - Chapter 10 – The Mole

Section 10.1 Measuring Matter

1. pair
2. 5
3. dozen
4. gross
5. 200
6. ream
7. 6,000,000,000
8. 0.5 mol
9. 6.02×10^{23}
10. four moles
11. $\frac{6.02 \times 10^{23} \text{ Cu atoms}}{1 \text{ mol Cu}}$
12. $\frac{1 \text{ mol CH}_4}{6.02 \times 10^{23} \text{ molecules CH}_4}$
13. $\frac{1 \text{ mol Xe}}{6.02 \times 10^{23} \text{ molecules Xe}}$
14. $\frac{6.02 \times 10^{23} \text{ molecules F}_2}{1 \text{ mol F}_2}$

Section 10.2 Mass and the Mole

1. false
2. true
3. false
4. true
5. true
6. true
7. false
8. true
9. b
10. d, a
11. a
12. b, c
13. d
14. c

Section 10.3 Moles of Compounds

1. carbon, 1 atom; hydrogen, 4 atoms
2. carbon, 1 atom; hydrogen, 1 atom; chlorine, 3 atoms
3. 1 mol C; 4 mol H
4. 1 mol C; 1 mol H; 3 mol Cl
5. 6.02×10^{23}
6. 1.81×10^{24}
7. 16.05 g/mol
8. 1.20×10^{24}

Section 10.4 Empirical and Molecular Formulas

1. The percent composition of a compound is the percent by mass of each of the elements in a compound.
2. Divide the mass of each element in the sample by the mass of the sample. Then multiply each quotient by 100.
3. c
4. c
5. b
6. b
7. c
8. c
9. d
10. $7.89 \text{ g K} \times 1 \text{ mol K} / 39.10 \text{ g K} = 0.202 \text{ mol K}$
 $2.42 \text{ g C} \times 1 \text{ mol C} / 12.01 \text{ g C} = 0.201 \text{ mol C}$
 $9.69 \text{ g O} \times 1 \text{ mol O} / 16.00 \text{ g O} = 0.606 \text{ mol O}$
 $0.202 \text{ mol K} / 0.201 \text{ mol C} = 1.00 \text{ mol K} / 1.00 \text{ mol C} = 1 \text{ mol K} / 1 \text{ mol C}$
 $0.201 \text{ mol C} / 0.201 \text{ mol C} = 1.00 \text{ mol C} / 1.00 \text{ mol C} = 1 \text{ mol C} / 1 \text{ mol C}$
 $0.606 \text{ mol O} / 0.201 \text{ mol C} = 3.01 \text{ mol O} / 1.00 \text{ mol C} = 3 \text{ mol O} / 1 \text{ mol C}$
1 mol K : 1 mol C : 3 mol O
empirical formula: KCO_3
 $(39.10 \text{ g} + 12.01 \text{ g} + 48.00 \text{ g}) = 99.11 \text{ g}$

Chapter 10 The Mole Study Guide Answers

Marco Cascella



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