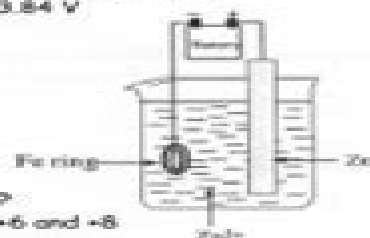


Electrochemistry - Multiple Choice Questions

- ___ 1. Which of these chemical reactions is an oxidation-reduction reaction?
 A. $\text{Fe} + \text{S} \rightarrow \text{FeS}$
 B. $\text{CO}_2 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{CO}_3$
 C. $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$
 D. $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$
- ___ 2. What happens to the oxidizing agent in an oxidation-reduction reaction?
 A. It is oxidized as it gains electrons.
 B. It is oxidized as it loses electrons.
 C. It is reduced as it gains electrons.
 D. It is reduced as it loses electrons.
- ___ 3. In which substance does bromine have an oxidation number of +1?
 A. Br_2
 B. HBr
 C. HBrO
 D. HBrO_2
- ___ 4. Which statement is true for an electrochemical cell?
 A. Oxidation occurs at the anode only.
 B. Reduction occurs at the anode only.
 C. Oxidation occurs at both the anode and cathode.
 D. Reduction occurs at both the anode and cathode.
- ___ 5. Given the equation: $2\text{Cr}(s) + 3\text{Pb}^{2+}(\text{aq}) \rightarrow 2\text{Cr}^{3+}(\text{aq}) + 3\text{Pb}(s)$, which is the correct reduction half reaction?
 A. $\text{Cr}(s) \rightarrow \text{Cr}^{3+}(\text{aq}) + 3\text{e}^-$
 B. $\text{Cr}(s) + 3\text{e}^- \rightarrow \text{Cr}^{3+}(\text{aq})$
 C. $\text{Pb}^{2+}(\text{aq}) \rightarrow \text{Pb}(s) + 2\text{e}^-$
 D. $\text{Pb}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Pb}(s)$
- ___ 6. What is the E° for an electrochemical cell with the following reaction? $2\text{Au}^{3+} + 3\text{Co} \rightarrow 3\text{Co}^{2+} + 2\text{Au}$
 A. -1.22 V
 B. 1.78 V
 C. 1.22 V
 D. 3.84 V
- ___ 7. An iron ring is plated with zinc metal in the apparatus below. Which of the following is true?
 A. It is a voltaic cell and the reaction is spontaneous.
 B. It is a voltaic cell and the reaction is not spontaneous.
 C. It is an electrolytic cell and the reaction is spontaneous.
 D. It is an electrolytic cell and the reaction is not spontaneous.
- ___ 8. What are the oxidation state of vanadium in the ions VO^{2+} and VO_4^{3-} respectively?
 A. +4 and +5
 B. +4 and +8
 C. +6 and +5
 D. +6 and +8
- ___ 9. Which one of the following reactions is a redox reaction?
 A. $\text{Pb}^{2+}(\text{aq}) + 2\text{Cl}^-(\text{aq}) \rightarrow \text{PbCl}_2(s)$
 B. $\text{AgNO}_3(\text{aq}) + \text{HCl}(\text{aq}) \rightarrow \text{AgCl}(s) + \text{HNO}_3(\text{aq})$
 C. $\text{NaOH}(\text{aq}) + \text{HCl}(\text{aq}) \rightarrow \text{NaCl}(\text{aq}) + \text{H}_2\text{O}(l)$
 D. $2\text{Al}(s) + 3\text{Cl}_2(g) \rightarrow 2\text{AlCl}_3(s)$
- ___ 10. Consider the following unbalanced redox equation:
 $\text{CH}_3\text{OH}(l) + \text{Cr}_2\text{O}_7^{2-}(\text{aq}) + \text{H}^+(\text{aq}) \rightarrow \text{CH}_2\text{O}(\text{aq}) + \text{Cr}^{3+}(\text{aq}) + \text{H}_2\text{O}(l)$
 Which of the following sets of numbers will balance the equation?
 A. 1, 1, 14, 1, 2, 7
 B. 3, 1, 8, 3, 2, 7
 C. 3, 1, 8, 3, 2, 6
 D. 3, 1, 14, 3, 2, 8
- ___ 11. In which of the following does sulfur have an oxidation number of +7?
 A. HSO_3^-
 B. SO_2
 C. H_2SO_4
 D. $\text{H}_2\text{S}_2\text{O}_8$
- ___ 12. What happens to the reducing agent in an oxidation-reduction reaction?
 A. It is oxidized as it gains electrons.
 B. It is oxidized as it loses electrons.
 C. It is reduced as it gains electrons.
 D. It is reduced as it loses electrons.
- ___ 13. What is the term for the electrode where oxidation occurs?
 A. anode
 B. cathode
 C. oxidizing agent
 D. reducing agent



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Answer

Fredeen, Pearson



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Table of Contents Chapter Electrochemistry Voltaic Cells Study Guide Answer

1. Understanding the eBook Chapter Electrochemistry Voltaic Cells Study Guide Answer
 - The Rise of Digital Reading Chapter Electrochemistry Voltaic Cells Study Guide Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter Electrochemistry Voltaic Cells Study Guide Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter Electrochemistry Voltaic Cells Study Guide Answer
 - User-Friendly Interface

4. Exploring eBook Recommendations from Chapter Electrochemistry Voltaic Cells Study Guide Answer
 - Personalized Recommendations
 - Chapter Electrochemistry Voltaic Cells Study Guide Answer User Reviews and Ratings
 - Chapter Electrochemistry Voltaic Cells Study Guide Answer and Bestseller Lists
5. Accessing Chapter Electrochemistry Voltaic Cells Study Guide Answer Free and Paid eBooks
 - Chapter Electrochemistry Voltaic Cells Study Guide Answer Public Domain eBooks
 - Chapter Electrochemistry Voltaic Cells Study Guide Answer eBook Subscription Services
 - Chapter Electrochemistry Voltaic Cells Study Guide Answer Budget-Friendly Options
6. Navigating Chapter Electrochemistry Voltaic Cells Study Guide Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter Electrochemistry Voltaic Cells Study Guide Answer Compatibility with Devices
 - Chapter Electrochemistry Voltaic Cells Study Guide Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter Electrochemistry Voltaic Cells Study Guide Answer
 - Highlighting and Note-Taking Chapter Electrochemistry Voltaic Cells Study Guide Answer
 - Interactive Elements Chapter Electrochemistry Voltaic Cells Study Guide Answer
8. Staying Engaged with Chapter Electrochemistry Voltaic Cells Study Guide Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter Electrochemistry Voltaic Cells Study Guide Answer
9. Balancing eBooks and Physical Books Chapter Electrochemistry Voltaic Cells Study Guide Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter Electrochemistry Voltaic Cells Study Guide Answer
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter Electrochemistry Voltaic Cells Study Guide Answer
 - Setting Reading Goals Chapter Electrochemistry Voltaic Cells Study Guide Answer
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chapter Electrochemistry Voltaic Cells Study Guide Answer
 - Fact-Checking eBook Content of Chapter Electrochemistry Voltaic Cells Study Guide Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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