



55. Phosphorus has three allotropic forms — (i) white phosphorus (ii) red phosphorus and (iii) black phosphorus. Write the difference between white and red phosphorus on the basis of their structure and reactivity.

Solution:

1. White phosphorus

(i) It is soft waxy solid having garlic order

(ii) It is poisonous

(iii) Has low melting point and boiling point because P4 molecules are held together by weak van der Waal forces of attraction

2. Red Phosphorus

(i) It is hard, crystalline, odourless solid

(ii) It is non-poisonous

(iii) It has a high melting point because of its polymeric structure.

56. Give an example to show the effect of concentration of nitric acid on the formation of oxidation product.

Solution:

4Zn +10HNO3(dil) 4Zn(NO3)2 + NH4NO3 + 3H2O

4Zn+10HNO3(dil) 4Zn(NO3)2+N2O+5H2O

4Zn +4HNO3(conc.) Zn(NO3)2 + NO2 + 2H2O

57. PCI5 reacts with finely divided silver on heating and a white silver salt is obtained, which dissolves on adding excess aqueous NH3 solution. Write the reactions involved to explain what happens.

Solution:

PCI5 reacts with silver to form white silver salt(AgCl). Which than dissolves on adding excess aqueous NH3 to form a soluble complex PCI5 + 2Ag 2AgCl(white ppt.) + PCI3 AgCl + 2NH3(aq) [Ag(NH3)]+Cl-

58. Phosphorus forms a number of oxoacids. Out of these oxoacidsphosphinicacid has strong

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Ralph M. Feather

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Table of Contents Chapter 17 Assessment Chemistry Glencoe Answers

- 1. Understanding the eBook Chapter 17 Assessment Chemistry Glencoe Answers
 - The Rise of Digital Reading Chapter 17 Assessment Chemistry Glencoe Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Chapter 17 Assessment Chemistry Glencoe Answers
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 17 Assessment Chemistry Glencoe Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chapter 17 Assessment Chemistry Glencoe Answers
 - Personalized Recommendations

- Chapter 17 Assessment Chemistry Glencoe Answers User Reviews and Ratings
- Chapter 17 Assessment Chemistry Glencoe Answers and Bestseller Lists
- 5. Accessing Chapter 17 Assessment Chemistry Glencoe Answers Free and Paid eBooks
 - Chapter 17 Assessment Chemistry Glencoe Answers Public Domain eBooks
 - Chapter 17 Assessment Chemistry Glencoe Answers eBook Subscription Services
 - Chapter 17 Assessment Chemistry Glencoe Answers Budget-Friendly Options
- 6. Navigating Chapter 17 Assessment Chemistry Glencoe Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Chapter 17 Assessment Chemistry Glencoe Answers Compatibility with Devices
 - Chapter 17 Assessment Chemistry Glencoe Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 17 Assessment Chemistry Glencoe Answers
 - Highlighting and Note-Taking Chapter 17 Assessment Chemistry Glencoe Answers
 - Interactive Elements Chapter 17 Assessment Chemistry Glencoe Answers
- 8. Staying Engaged with Chapter 17 Assessment Chemistry Glencoe Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 17 Assessment Chemistry Glencoe Answers
- 9. Balancing eBooks and Physical Books Chapter 17 Assessment Chemistry Glencoe Answers
 - ∘ Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 17 Assessment Chemistry Glencoe Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 17 Assessment Chemistry Glencoe Answers
 - Setting Reading Goals Chapter 17 Assessment Chemistry Glencoe Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 17 Assessment Chemistry Glencoe Answers
 - Fact-Checking eBook Content of Chapter 17 Assessment Chemistry Glencoe Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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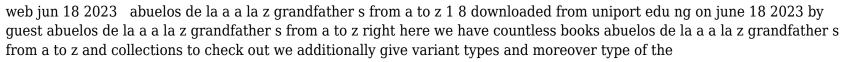
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