

## Reviewing Content

42. The solvent is the substance in which the solute is dissolved.
43. Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
44. **solubility**: the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.  
**saturated solution**: a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure. **unsaturated solution**: a solution that contains less solute than a saturated solution at a given temperature and pressure. **miscible**: describes liquids that dissolve in each other. **immiscible**: describes liquids that are insoluble in each other.
45. Particles of solute crystallize.
46. No; if there were undissolved solute, the excess solute would come out of a supersaturated solution.
47.  $5.55 \times 10^2$  g  $\text{AgNO}_3$
48. Solubility increases with pressure.
49. a.  $1.6 \times 10^{-2}$  g/L  
b.  $4.7 \times 10^{-2}$  g/L
50. *Dilute* and *concentrated* are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
51. Molarity is the number of moles of solute dissolved in one liter of solution.  
a. 1.3M KCl  
b.  $3.3 \times 10^{-1}$ M  $\text{MgCl}_2$
52.  $2.00 \times 10^3$  mL
53. a.  $5.0 \times 10^{-3}$  mol NaCl, 29 g NaCl  
b. 1.0 mol  $\text{KNO}_3$ ,  $1.0 \times 10^{-2}$  g  $\text{KNO}_3$   
c.  $2.5 \times 10^3$  mol  $\text{CaCl}_2$ , 2.8 g  $\text{CaCl}_2$
54. a.  $2.3 \times 10^1$  g NaCl  
b. 2.0 g  $\text{MgCl}_2$
55. a. 16% (v/v) ethanol  
b. 63.6% (v/v) isopropyl alcohol
56. Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation, freezing-point depression, and vapor-pressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution; relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
57. a. sea water  
b. 1.5M  $\text{KNO}_3$   
c. 0.100M  $\text{MgCl}_2$
58. The effective molality of the  $\text{Ca}(\text{NO}_3)_2$  solution is 3m. The effective molality of the  $\text{NaNO}_3$  solution is 2m.
59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
60. The salt lowers the freezing point of the ice-water cooling mixture.
61. 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
62. Add 27.0 g  $\text{H}_2\text{O}$  to 32.0 g  $\text{CH}_3\text{OH}$ .
63. a.  $100.26^\circ\text{C}$   
b.  $101.54^\circ\text{C}$
64. a.  $-4.46^\circ\text{C}$   
b.  $-2.2^\circ\text{C}$
65. a.  $-1.1^\circ\text{C}$   
b.  $-0.74^\circ\text{C}$   
c.  $-1.5^\circ\text{C}$

## Understanding Concepts

66. a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

# Chapter 16 Chemistry Study Guide Answers

**Richard Daley**



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## **Table of Contents Chapter 16 Chemistry Study Guide Answers**

1. Understanding the eBook Chapter 16 Chemistry Study Guide Answers
  - The Rise of Digital Reading Chapter 16 Chemistry Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 16 Chemistry Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 16 Chemistry Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 16 Chemistry Study Guide Answers
  - Personalized Recommendations
  - Chapter 16 Chemistry Study Guide Answers User Reviews and Ratings
  - Chapter 16 Chemistry Study Guide Answers and Bestseller Lists
5. Accessing Chapter 16 Chemistry Study Guide Answers Free and Paid eBooks
  - Chapter 16 Chemistry Study Guide Answers Public Domain eBooks
  - Chapter 16 Chemistry Study Guide Answers eBook Subscription Services
  - Chapter 16 Chemistry Study Guide Answers Budget-Friendly Options

6. Navigating Chapter 16 Chemistry Study Guide Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 16 Chemistry Study Guide Answers Compatibility with Devices
  - Chapter 16 Chemistry Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 16 Chemistry Study Guide Answers
  - Highlighting and Note-Taking Chapter 16 Chemistry Study Guide Answers
  - Interactive Elements Chapter 16 Chemistry Study Guide Answers
8. Staying Engaged with Chapter 16 Chemistry Study Guide Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 16 Chemistry Study Guide Answers
9. Balancing eBooks and Physical Books Chapter 16 Chemistry Study Guide Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 16 Chemistry Study Guide Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 16 Chemistry Study Guide Answers
  - Setting Reading Goals Chapter 16 Chemistry Study Guide Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 16 Chemistry Study Guide Answers
  - Fact-Checking eBook Content of Chapter 16 Chemistry Study Guide Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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