Answer Key

Chapter 14

Practice Problems

- a. 61 g CaCl₂/100 g H₂O
 b. 90 g CaCl₂/100 g H₂O
 - c. 5 g Ce₂(SO₄)₃
 - d. 1 g Ce₂(SO₄)₃
 - e. 60 g KClO₃
- 2. 94 atm
- 3. 2.1 atm
- 4. 0.852%
- 5. 7.35 g
- 6. 492.65 g
- 7. 26%
- 8. 0.868 L
- 9. 0.328M
- 10. 1.09M
- 11. 326 mL

- 12. 0.802M
- 13. 0.350m
- 14. 0.0300m
- a. water, 0.830; ethanol, 0.170
 - b. water, 0.9853; calcium chloride, 0.0147
 - c. carbon tetrachloride,
 0.644; benzene, 0.356
- 16. BP = 102.06°C; FP = -7.48°C
- 17. 2.33m
- 18. a. suspension
 - b. solution or colloid
 - c. colloid
 - d. solution

Chapter 14 Review

- 19. Solvation is a process in which solvent particles surround solute particles to form a solution. A substance dissolves when attractive forces between solvent and solute particles are strong enough to overcome the attractive forces holding the solute particles together.
- 20. A saturated solution contains the maximum amount of dissolved solute for a given amount of solvent at a specific temperature and pressure. An unsaturated solution contains less dissolved solute for a given temperature and pressure than does a saturated solution. A supersaturated solution contains more dissolved solute than does a saturated solution at the same temperature.

Chapter 14 Study Guide Mixtures Solutions Answers

Neil E. Schore

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