

**10****CHEMICAL QUANTITIES****Practice Problems**

*On a piece of paper, solve the following problems.*

***YOU MUST SHOW YOUR WORK TO GET CREDIT FOR YOUR ANSWERS!!!***

**SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER**

1. What is the molar mass of sucrose ( $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ )?
2. What is the molar mass of each of the following compounds?
  - a. phosphorus pentachloride ( $\text{PCl}_5$ )
  - b. uranium hexafluoride ( $\text{UF}_6$ )
3. Calculate the molar mass of each of the following ionic compounds:
  - a.  $\text{KMnO}_4$
  - b.  $\text{Ca}_3(\text{PO}_4)_2$
4. How many moles is  $3.52 \times 10^{24}$  molecules of water?
5. How many atoms of zinc are in 0.60 mol of zinc?
6. What is the mass of 1.00 mol of oxygen ( $\text{O}_2$ )?

**SECTION 10.2 MOLE–MASS AND MOLE–VOLUME RELATIONSHIPS**

1. What is the molar mass of each of the following compounds?
  - a.  $\text{C}_6\text{H}_{12}\text{O}_6$
  - b.  $\text{NaHCO}_3$
  - c.  $\text{C}_3\text{H}_{12}$
  - d.  $\text{KNH}_4\text{SO}_4$
2. Calculate the mass in grams of each of the following:
  - a. 8.0 mol lead oxide ( $\text{PbO}$ )
  - b. 0.75 mol hydrogen sulfide ( $\text{H}_2\text{S}$ )
  - c. 0.00100 mol silicon tetrahydride ( $\text{SiH}_4$ )
  - d.  $1.50 \times 10^{-2}$  mol molecular oxygen ( $\text{O}_2$ )
  - e. 2.30 mol ethylene glycol ( $\text{C}_2\text{H}_6\text{O}_2$ )
3. How many grams are in 1.73 mol of dinitrogen pentoxide ( $\text{N}_2\text{O}_5$ )?
4. How many grams are in 0.658 mol of calcium phosphate [ $\text{Ca}_3(\text{PO}_4)_2$ ]?
5. Calculate the number of moles in each of the following:
  - a. 0.50 g sodium bromide ( $\text{NaBr}$ )
  - b. 13.5 g magnesium nitrate [ $\text{Mg}(\text{NO}_3)_2$ ]
  - c. 1.02 g magnesium chloride ( $\text{MgCl}_2$ )
  - d. 0.00100 g monochloromethane ( $\text{CH}_3\text{Cl}$ )
  - e.  $1.50 \times 10^{-3}$  g propylene glycol [ $\text{C}_3\text{H}_8(\text{OH})_2$ ]
6. A chemist plans to use 435.0 grams of ammonium nitrate ( $\text{NH}_4\text{NO}_3$ ) in a reaction. How many moles of the compound is this?
7. A solution is to be prepared in a laboratory. The solution requires 0.0465 mol of quinine ( $\text{C}_{20}\text{H}_{24}\text{N}_2\text{O}_2$ ). What mass, in grams, should the laboratory technician obtain in order to make the solution?
8. What is the volume at STP of 2.66 mol of methane ( $\text{CH}_4$ ) gas?
9. How many moles is 135 L of ammonia ( $\text{NH}_3$ ) gas at STP?

# Chapter 10 Study Guide Chemical Reactions Answers

**Barron's Educational Series, Mark  
Kernion, Joseph A. Mascetta**



## **Chapter 10 Study Guide Chemical Reactions Answers:**

**Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone, Theodore O. Dolter, 2012-01-03

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**Study Guide/Selected Solutions Manual** Julie R. Frentrop, 2002-06. Contains a brief overview of every chapter, review of skills, self-tests, and the answers and detailed solutions to all odd-numbered end-of-chapter problems in the text book.

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**Study Guide with Answers to Selected Problems** H. Stephen Stoker, Danny V. White, 2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

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make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students *Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today* Seager/Slabaugh,2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application *Chemistry, Student Study Guide* James E. Brady,Fred Senese,2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement [Class 9 Physics Questions and Answers PDF](#) Arshad Iqbal, The Class 9 Physics Quiz Questions and Answers PDF Grade 9 Physics Competitive Exam Questions Chapter 1 9 Practice Tests Class 9 Physics Textbook Questions for Beginners includes revision guide for problem solving with hundreds of solved questions Class 9 Physics Questions and Answers PDF book covers basic concepts analytical and practical assessment tests Class 9 Physics Quiz PDF book helps to practice test questions from exam prep notes The Grade 9 Physics Quiz Questions and Answers PDF eBook includes revision guide with verbal quantitative and analytical past papers solved tests Class 9 Physics Questions and Answers PDF Free download chapter 1 a book covers solved common questions and answers on chapters Dynamics gravitation kinematics matter properties physical quantities and measurement thermal properties of matter transfer of heat turning effect of forces work and energy tests for school and college revision guide Physics Interview Questions and Answers PDF Download free eBook s sample covers beginner s solved questions textbook s study notes to practice online tests The Class 9 Physics Interview Questions Chapter 1 9 PDF book includes high school question papers to review practice tests for exams Class 9 Physics Practice Tests a textbook s revision guide with chapters

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## **Table of Contents Chapter 10 Study Guide Chemical Reactions Answers**

1. Understanding the eBook Chapter 10 Study Guide Chemical Reactions Answers
  - The Rise of Digital Reading Chapter 10 Study Guide Chemical Reactions Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 10 Study Guide Chemical Reactions Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 10 Study Guide Chemical Reactions Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 10 Study Guide Chemical Reactions Answers
  - Personalized Recommendations
  - Chapter 10 Study Guide Chemical Reactions Answers User Reviews and Ratings

- Chapter 10 Study Guide Chemical Reactions Answers and Bestseller Lists
- 5. Accessing Chapter 10 Study Guide Chemical Reactions Answers Free and Paid eBooks
  - Chapter 10 Study Guide Chemical Reactions Answers Public Domain eBooks
  - Chapter 10 Study Guide Chemical Reactions Answers eBook Subscription Services
  - Chapter 10 Study Guide Chemical Reactions Answers Budget-Friendly Options
- 6. Navigating Chapter 10 Study Guide Chemical Reactions Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 10 Study Guide Chemical Reactions Answers Compatibility with Devices
  - Chapter 10 Study Guide Chemical Reactions Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 10 Study Guide Chemical Reactions Answers
  - Highlighting and Note-Taking Chapter 10 Study Guide Chemical Reactions Answers
  - Interactive Elements Chapter 10 Study Guide Chemical Reactions Answers
- 8. Staying Engaged with Chapter 10 Study Guide Chemical Reactions Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 10 Study Guide Chemical Reactions Answers
- 9. Balancing eBooks and Physical Books Chapter 10 Study Guide Chemical Reactions Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 10 Study Guide Chemical Reactions Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 10 Study Guide Chemical Reactions Answers
  - Setting Reading Goals Chapter 10 Study Guide Chemical Reactions Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 10 Study Guide Chemical Reactions Answers
  - Fact-Checking eBook Content of Chapter 10 Study Guide Chemical Reactions Answers
  - Distinguishing Credible Sources



13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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