

Chemical Reactions

Question 1: Reactions and Equations

For each statement, write an equation or chemical equation to represent it. If it is not possible to write an equation, write "Not possible".

1. A reaction occurs between two substances.
2. A reaction occurs between two substances and a gas is produced.
3. A reaction occurs between two substances and a solid is produced.
4. A reaction occurs between two substances and a gas is produced.
5. A reaction occurs between two substances and a solid is produced.
6. A reaction occurs between two substances and a gas is produced.
7. A reaction occurs between two substances and a solid is produced.
8. A reaction occurs between two substances and a gas is produced.

Write the balanced chemical equation for each reaction. Write the names of the reactants, products, and any catalysts.

Reactants	Products	Equation	Notes
Hydrogen	Hydrogen	$2H_2 \rightarrow 2H_2$	Not a reaction

Write the balanced chemical equation for each reaction. Write the names of the reactants, products, and any catalysts.

1. Hydrogen reacts with oxygen to form water. $2H_2 + O_2 \rightarrow 2H_2O$

2. Hydrogen reacts with oxygen to form water. $2H_2 + O_2 \rightarrow 2H_2O$

3. Hydrogen reacts with oxygen to form water. $2H_2 + O_2 \rightarrow 2H_2O$

4. Hydrogen reacts with oxygen to form water. $2H_2 + O_2 \rightarrow 2H_2O$

5. Hydrogen reacts with oxygen to form water. $2H_2 + O_2 \rightarrow 2H_2O$

6. Hydrogen reacts with oxygen to form water. $2H_2 + O_2 \rightarrow 2H_2O$

7. Hydrogen reacts with oxygen to form water.

Chapter 10 Chemical Reactions Study Guide For Content Mastery

Julia Schneider



Chapter 10 Chemical Reactions Study Guide For Content Mastery:

Reviewing **Chapter 10 Chemical Reactions Study Guide For Content Mastery**: Unlocking the Spellbinding Force of Linguistics

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