

Chemistry – Unit 5 Worksheet 3

Empirical and Molecular Formulas

Show all your work when solving the following problems. Be sure to include units and label your answer.

1. Find the empirical formula of a compound containing 32.0 g of bromine and 4.9 g of magnesium.

$$32.0 \text{ g} \times \frac{1 \text{ mole}}{79.9 \text{ g}} = 0.401 \text{ mole Br}$$

1:2 ratio

$$4.9 \text{ g} \times \frac{1 \text{ mole}}{24.31 \text{ g}} = 0.20 \text{ mole Mg}$$

 MgBr_2

2. What is the empirical formula of a carbon-oxygen compound, given that a 95.2 g sample of the compound contains 40.8 g of carbon and the rest oxygen?

$$95.2 \text{ g} - 40.8 \text{ g} = 54.4 \text{ g O}$$

$$40.8 \text{ g} \times \frac{1 \text{ mole}}{12.0 \text{ g}} = 3.40 \text{ moles C}$$

1:1 ratio

$$54.4 \text{ g} \times \frac{1 \text{ mole}}{16.0 \text{ g}} = 3.40 \text{ moles O}$$

 CO

3. A compound was analyzed and found to contain 9.8 g of nitrogen, 0.70 g of hydrogen, and 33.6 g of oxygen. What is the empirical formula of the compound?

$$9.8 \text{ g} \times \frac{1 \text{ mole}}{14.0 \text{ g}} = 0.70 \text{ moles N}$$

1:1:3 ratio

$$0.70 \text{ g} \times \frac{1 \text{ mole}}{1.01 \text{ g}} = 0.69 \text{ moles H}$$

 HNO_3

$$33.6 \text{ g} \times \frac{1 \text{ mole}}{16.0 \text{ g}} = 2.10 \text{ moles O}$$

4. A compound composed of hydrogen and oxygen is found to contain 0.59 g of hydrogen and 9.40 g of oxygen. The molar mass of this compound is 34.0 g/mol. Find the empirical and molecular formulas.

$$0.59 \text{ g} \times \frac{1 \text{ mole}}{1.01 \text{ g}} = 0.58 \text{ mole H}$$

1:1 ratio; empirical formula: HO
 empirical formula mass: 17.0 g/mol
 molecular formula mass: 34.0 g/mol
 mfm is 2x efm so the molecular formula is: H_2O_2

$$9.40 \text{ g} \times \frac{1 \text{ mole}}{16.0 \text{ g}} = 0.588 \text{ mole O}$$

Empirical Formula Answer Sheet

RJ Shavelson

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Empirical Formula Answer Sheet:

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Clifford C. Houk, Richard Post, 1996-03-09 CHEMISTRY SECOND EDITION The fast easy way to master the fundamentals of chemistry Have you ever wondered about the differences between liquids gases and solids Or what actually happens when something burns What exactly is a solution An acid A base This is chemistry the composition and structure of substances composing all matter and how they can be transformed Whether you are studying chemistry for the first time on your own want to refresh your memory for a test or need a little help for a course this concise interactive guide gives you a fresh approach to this fascinating subject This fully up to date edition of Chemistry Concepts and Problems Has been tested rewritten and retested to ensure that you can teach yourself all about chemistry Requires no prerequisites Lets you work at your own pace with a helpful question and answer format Lists objectives for each chapter you can skip ahead or find extra help if you need it Reinforces what you learn with chapter self tests **Regents Chemistry--Physical Setting Power Pack**

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