

## Empirical Formulas and Molecular Formulas

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### Activity 1

Remember, we already completed an empirical formula calculation using the mass percentages for oxygen and hydrogen for a compound. Do you recall the empirical formula of the compound that we used to model different types of polymers? What were the assumptions we began with based on?

### Table of the Chemical Data of Observed Compounds and Empirical Formulas

Compound	Empirical Formula	Observed Molar Mass	Calculated Molar Mass	Empirical Composition Ratio	Percent Composition
1	CH <sub>2</sub>	14.03	14.03		
2	CH <sub>2</sub>	14.03	14.03		
3	CH <sub>2</sub>	14.03	14.03		
4	CH <sub>2</sub>	14.03	14.03		
5	CH <sub>2</sub>	14.03	14.03		

- Circle up the data table and calculate the percent composition for each of the polymer models. Do the values from the table above (and calculated) follow?
- Identify the polymer models in Model 1 that have the same empirical formula.
- Identify the polymers in Model 1 that have the same percent composition.
- Compare the polymer in Model 1, polyethylene, to the other polymers with different molecular formulas. Do the other polymers have the same percent composition?
- Identify polymers in Model 1 that all have the percent formula calculated in Model 1 except the units. Polyethylene is CH<sub>2</sub> and the other models are CH<sub>2</sub>. Do you see the empirical formula of polyethylene in the polymer's molecular formula?

# Empirical Formula And Molecular Formula Pogil

**Sebastian Brünink**



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