2)	Based upon the graph in question 1, how does a decrease in temperature affect the volume of the gas (the pressure and amount of gas is held constant)?
	Answer:
3)	Charles' Law allows the prediction of gas volumes at different temperatures. What would the volume of a gas be at 38.3°C if it started at a volume of 45.7mL and a temperature of -58.5 °C? Remember that temperatures must be in Kelvin when performing Gas Law calculations.
	Answer:
4)	Using the raw data from question 3 as an example, explain why temperature must always be converted to Kelvin when using gas law equations.
	Answer:

Charles Law Pre Lab Answers

Victor M. Corman

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