

## Lab Data



Verify your mass calculation

Verify your molality calculation

Did you account for the freezing-point of your DI water?

Did you report your data to the correct number of significant figures?

	Solution 1	Solution 2	Solution 3
Magnesium chloride	MgCl <sub>2</sub>	MgCl <sub>2</sub>	MgCl <sub>2</sub>
Molar mass (g/mol)	95.21	95.21	95.21
Mass of calorimeter (g)	17.22		
Volume of DI water (mL)	47.1		
Mass of magnesium chloride (g)	1.58		
Moles of magnesium chloride (mol)	0.0166		
Mass of calorimeter + solution (g)	66.12		
Mass of calorimeter + solution + ice (g)	110.45		
Freezing-point of solution (°C)	-1.42		
Mass of calorimeter + ice after decanting (g)	39.88		
Mass of ice added (g)	44.33		
Mass of ice remaining after decanting (g)	22.66		
Mass of melted ice (g)	21.67		
Mass of water in final solution (g)	<input type="text"/>		
Mass of water in final solution (kg)	<input type="text"/>		
Molality of solution (m)	<input type="text"/>		
Freezing-point depression (°C)	<input type="text"/>		

van't Hoff factor

# Colligative Properties Virtual Lab Answers Key

**S Marginson**



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