Last Name (Surname)

1. (16 points) Using the reduction potentials provided, circle the appropriate choice in each case.

See explanation below

$$VO_3^+ + 2H^+ + e^- \rightarrow VO^{3+} + H_3O$$

$$\epsilon^{\circ} = 0.99 \text{ V}$$

(circle one)

$$E^{\circ} = 1.33 \text{ V}$$

Strongest Oxidizing Agent: (

Strongest Reducing Agent:

$$\epsilon^* = -1.18 \text{ V}$$

$$AI^{2*} + 3e^* \Rightarrow AI$$

Strongest Reducing Agent:

(circle one)

$$E^{\circ} = -0.73 \text{ V}$$

$$Zn^{2+} + 2e^{-} \rightarrow Zn$$

$$E^* = -0.76 \text{ V}$$

Strongest Oxidizing Agent:



(circle one)

$$Au^{3+} + 3e^- \rightarrow Au$$

$$E^{\circ} = 1.50 \text{ V}$$

$$Ag^* + e^* \rightarrow Ag$$

[Fe(NO,),]*

(6 points) Determine the oxidation number in each case.

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