

Sample Placement Exam for Chemistry 200

The Placement Exam contains 44 multiple-choice questions. You will be allowed 45 minutes to do the exam. A good workbook for reviewing for the test is "Schaum's Easy Outlines, Beginning Chemistry". Silent, non-printing, non-programmable calculators will be permitted; Personal Digital Assistants will not be permitted. Your score will be the total of correct answers. A periodic table and other useful information will be provided on the back page. Scratch paper will also be provided.

The exam covers the following topics in chemistry and mathematics at the high school level.
Elements and Atoms: basic properties, atomic number, atomic mass, isotopes, Lewis symbols, valence electrons, electronic configurations, use of the periodic table

Compounds and Molecules: physical properties, molecular mass, naming inorganic compounds, formula calculations, simple Lewis dot structures

Reactions: Balancing equations, simple stoichiometry, equilibria, enthalpy, entropy, simple kinetics

Bonding: Ionic, covalent, metallic

Gases: Use of the ideal gas law

Solutions: Molarity, dilution, basic ideas of acids and bases, properties of electrolytes

Mathematics: significant figures, simple algebra

The following twelve questions are examples of the types of questions that may be asked, but note that they do not cover all of the material that may be on the exam.

Circle the one best answer for each of the following questions.

- What is the molar mass (molecular weight) of the compound Al_2O_3 ?
A. 21.00
B. 42.98
C. 50.00
D. 101.96
- The density of ethanol is 0.789 g/cm^3 . How many grams are there in 205 milliliters of ethanol?
A. 162
B. 205
C. 260
D. 2.60×10^3
- Which chemical equation below is *balanced*?
A. $2 \text{N}_2(\text{g}) + 3 \text{H}_2(\text{g}) \rightarrow 2 \text{NH}_3(\text{g})$
B. $2 \text{Mg}(\text{s}) + \text{O}_2(\text{g}) \rightarrow 2 \text{MgO}(\text{s})$
C. $2 \text{Al}(\text{s}) + 3 \text{O}_2(\text{g}) \rightarrow \text{Al}_2\text{O}_3(\text{s})$
D. $\text{Ca}[\text{OH}]_2(\text{s}) + \text{CO}_2(\text{g}) \rightarrow \text{Ca}[\text{HCO}_3]_2(\text{s})$
- How many grams of sugar are in six dozen doughnuts if each doughnut contains 75 g of sugar?
A. 0.96
B. 450
C. 900
D. 5400

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