1. How much heat will be released when 4.72 g of carbon react with excess O<sub>2</sub>?

$$C + O_2 \rightarrow CO_2$$
;  $\Delta H^\circ = -393.5 \text{ kJ}$ 

2. How much heat will be absorbed when 38.2 g of bromine react with excess H<sub>2</sub>?

$$H_2 + Br_2 \rightarrow 2 HBr; \Delta H^\circ = +72.80 kJ$$

3. How much heat will be released when 1.48 g of chlorine react with excess phosphorus?

4. How much heat will be absorbed when 13.7 g of nitrogen react with excess O<sub>2</sub>?

$$N_2 + O_2 \rightarrow 2 \text{ NO}; \Delta H^\circ = +180 \text{ kJ}$$

Find the heat of reaction for the single displacement reaction in which 2 L of chlorine gas at STP react with sodium bromide.

$$Cl_2(g) + 2 \text{ NaBr (aq)} \rightarrow Br_2(l) + 2 \text{ NaCl (aq)}; \Delta H^\circ = -100.2 \text{ kJ}$$

6. A chemist runs an experiment where the following reaction occurs:

$$CrO_3$$
 (s) +  $H_2O$  (l)  $\rightarrow$   $H_2CrO_4$  (aq);  $\Delta H^\circ$  = -5.4 kJ

Careful measurement indicates that  $6.18 \, kJ$  of energy were released. What mass of  $CrO_3$  reacted?

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**Patrick Vollmar** 

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