



Name _____ Class _____ Date _____

1

How many atoms of **oxygen (O)** are in the **chemical formula** below?

- A** 4
- B** 5
- C** 7
- D** 12

**3**

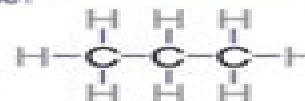
Carbon (C) and hydrogen (H) create **covalent bonds** to share **electrons** between them. What is the total number of electrons being shared between **carbon** and **hydrogen** in the molecule of propane shown below?

- A** 8
- B** 11
- C** 16
- D** 32

**2**

The diagram shown below is a **structural formula** for the gas **propane**. What would be the correct **chemical formula** for propane?

- A** C_3H_6
- B** CH_6
- C** $8\text{H}_3\text{C}$
- D** CH

**4**

The **reason** why the bonds between carbon (C) and hydrogen (H) are **covalent** is _____

- A** C and H do not have electrons to give away
- B** C and H are both nonmetals
- C** C and H are both metalloids
- D** C and H can bond only by sharing electrons

**5**

The diagram below represents the **structural formula** for several **water molecules**. How many **molecules** are represented?

- A** 3
- B** 5
- C** 10
- D** 15

**6**

Explain the **arrangement** of the molecules in the diagram of water below.

- A** it is a lucky arrangement
- B** water molecules are sticky
- C** water molecules have polarity
- D** water molecules are magnetic

**7**

Water molecules have some **polarity** in their structure. Using the diagram below, notice how the ends of H atoms (**positive**) and O atoms (**negative**) meet. This property is the **reason** that _____

- A** water forms droplets easily
- B** water boils easily
- C** ice melts easily
- D** water has color

**8**

Since H atoms are slightly **positive** and the O atoms are slightly **negative**, molecules of water have **polarity**. This makes water a **good** _____

- A** solute
- B** refrigerant
- C** solvent
- D** magnet

**9**

The chemical formula for the **oxygen** that we breathe is _____

- A** O
- B** O_2
- C** O_3
- D** 2O

**10**

How many **molecules of oxygen** are represented in the chemical formula below?

- A** 2
- B** 4
- C** 6
- D** 18



Fun Eighth Grade Science Questions

Katrin Zwirgmaier



Fun Eighth Grade Science Questions:

This book delves into Fun Eighth Grade Science Questions. Fun Eighth Grade Science Questions is a crucial topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Fun Eighth Grade Science Questions, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Fun Eighth Grade Science Questions
 - Chapter 2: Essential Elements of Fun Eighth Grade Science Questions
 - Chapter 3: Fun Eighth Grade Science Questions in Everyday Life
 - Chapter 4: Fun Eighth Grade Science Questions in Specific Contexts
 - Chapter 5: Conclusion
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 3. In chapter 2, the author will delve into the foundational concepts of Fun Eighth Grade Science Questions. This chapter will elucidate the essential principles that need to be understood to grasp Fun Eighth Grade Science Questions in its entirety.
 4. In chapter 3, the author will examine the practical applications of Fun Eighth Grade Science Questions in daily life. This chapter will showcase real-world examples of how Fun Eighth Grade Science Questions can be effectively utilized in everyday scenarios.
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 6. In chapter 5, this book will draw a conclusion about Fun Eighth Grade Science Questions. The final chapter will summarize the key points that have been discussed throughout the book.
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