

Chapter 1

- Chemistry:** the science that tries to understand how matter behaves by studying how and why it reacts with other
- experiments:** having controlled observations designed to test or disprove hypotheses
- hypothesis:** a testable interpretation or explanation of the observation
- scientific method:** a way of learning that emphasizes observation and experimentation
- scientific law:** a brief statement that summarizes past observations
- scientific theory:** a broader and deeper explanation for observations and laws

What is the difference between scientific laws and theories?
 Scientific laws have been proved many times and no exceptions
 Scientific theories had a broad explanation and lots of evidence
What are the steps of the scientific method?
 1. observation
 2. hypothesis
 3. experiment
 4. form laws / theories

What is the law of conservation of mass? by Antoine Lavoisier
 matter is neither created nor destroyed

What is Atomic theory?

all matter is composed of small, indivisible particles called atoms
 Building chemistry requires correctness, consistency, common sense

Chapter 2

What is a derived unit?
 represents the position of a measured quantity
Why are units important?
 so we know what the numbers mean / represent

Why are sig figs important?

- to let others know or gain precision during mathematical operation
- When are they useful?
 - all non-zero digits are significant
 - zeros between two numbers are significant
 - leading zeros that follow after a decimal point are significant
 - leading zeros before a decimal point are significant
 - leading zeros are not significant
- trailing zeros at the end of a number but before an implied decimal point are ambiguous and should be avoided
- practice (how many sig. figs?)
 - a. 23.45
 - b. 240.0 cm
 - c. 0.0456 mL
 - d. 450,789 cm
- answer key:
 - a. 4
 - b. 4
 - c. 4
 - d. 6

What is density? Equation?
 mass / volume

- practice
 - a. What is the density of a block of metal that occupies 200 cm³ and a mass of 742 g?

Conversion

What is a conversion factor?

- the relationship between units that is used to change the units of a measured quantity without changing the value
- use fact we know to help understanding easier (1 in = 2.54 cm)
- practice
 - a. 23.5 in to feet
 - b. 1.25 gal to liters

Reminder: Scientific notation is the g x 10ⁿ

Chapter 3

do the properties change of an element change when combined with another?

Each compound has its own properties

The law of constant composition (discovered by Proust) is based on a certain compound has the same proportions of its elements in a chemical formula. the mass always goes first.

molecular formula: the actual number of atoms of each element in a molecule of the compound. (C₂H₄O₂)

empirical formula: the simplest whole-number ratio of atoms of each element in a compound. (CH₂O)

atomic symbols have single letters as their basic units
 molecular elements do not normally start with single letters

- diatomic molecules - two atoms of that element bonded together
 - a. H, N, O, Cl, Br, I, S, F
- common element
- subscript means
 - forming one compound:
 - metal and one or more nonmetals
 - with binary ionic compounds (metal + nonmetal) you add "ide" to the end
 - when forming cation, use roman numerals to show the charge
- naming molecular compounds
 - two (or more) nonmetals
 - prefix, root, -ide

Chapter 3

What is matter?

anything that occupies space and has mass

What are the three states of matter? (and their properties?)

- solid - atoms + molecules are packed close + have fixed positions
- crystalline - atoms + molecules arranged in geometric patterns
- amorphous - atoms + molecules don't have long-range order
- liquid - atoms + molecules are close, move freely, but change shape
- gas - atoms + molecules are separated and move freely

What is a mixture?

composed of 2+ different types of atoms or molecules combined

What are the two types? (and properties?)

- a. heterogeneous: different composition
- b. homogeneous: same composition throughout (uniform)

What is a pure substance? element? compound?

- pure substance: composed of one or more of an element or molecule
- element: a substance that cannot be broken down
- compound: a substance composed of 2+ elements in fixed proportions

What is the difference between a mixture and a compound?

Know just look at definition

physical property: a property displayed without changing its composition
chemical property: a property displayed only through a composition change
physical change: a change with the composition does not change
chemical change: a change with the composition does change

Energy: the capacity to do work

kinetic energy: energy associated with motion

potential energy: energy associated with position or composition

What are the three common units for energy?

Joules, cal, kJ

exothermic: releases energy

endothermic: absorbs energy

What is heat capacity?

the number of heat units needed to raise the temperature

Classifying matter is important.

Find your own practice problems if needed

Temperature is heat

Chapter 4

Demonstrates

What he did first person to record the idea of atoms

a. matter is composed of small, indivisible particles

Thomson

discovered the electron

Rutherford

did his experiment

the penetrating experiment

a. proved it wrong (by Thomson)

Mendeleev

was able to find the pattern with elements

Bohr

atomic theory

a. each element is composed of tiny indivisible particles (atoms)

b. atoms of each element have the same mass + other properties

c. atoms combine in simple whole number ratios

What are the main ideas in the nuclear theory of an atom?

protons + neutrons, located in the middle, electrons + other small particles

subatomic particles:

1. electron - negative charge

2. proton - positive charge

3. neutron - neutral charge

electrical charge is a fundamental property

What is matter? is usually neutral

periodic table is when elements are arranged in order of increasing mass

the table is organized by increasing atomic #

metals

properties: good conductors, malleable, shiny

where: left side of the periodic table

nonmetals

properties: bad conductors, poor conductors

where: upper right side, negative charge

metalloids

properties: mixed properties, semiconductors

where: right side and top

What is a family or group?

a thing of elements with similar properties

know where alkali metals, alkaline earth metals, halogens, noble gases

are on the periodic table

ion: atom that loses or gains electrons

cation: positive charge

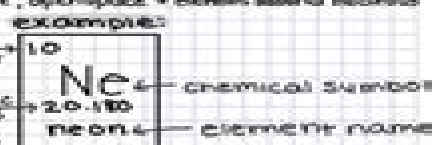
anion: negative charge

isotopes:

percent natural abundance: how much an element appears in nature

mass number: protons + neutrons

the number of electrons changes



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Table of Contents Ccsd Chemistry 1 Study Guide Semester

1. Understanding the eBook Ccsd Chemistry 1 Study Guide Semester
 - The Rise of Digital Reading Ccsd Chemistry 1 Study Guide Semester
 - Advantages of eBooks Over Traditional Books
2. Identifying Ccsd Chemistry 1 Study Guide Semester
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ccsd Chemistry 1 Study Guide Semester
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ccsd Chemistry 1 Study Guide Semester
 - Personalized Recommendations
 - Ccsd Chemistry 1 Study Guide Semester User Reviews and Ratings
 - Ccsd Chemistry 1 Study Guide Semester and Bestseller Lists
5. Accessing Ccsd Chemistry 1 Study Guide Semester Free and Paid eBooks
 - Ccsd Chemistry 1 Study Guide Semester Public Domain eBooks
 - Ccsd Chemistry 1 Study Guide Semester eBook Subscription Services
 - Ccsd Chemistry 1 Study Guide Semester Budget-Friendly Options
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-
- ePub, PDF, MOBI, and More
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 - Ccsd Chemistry 1 Study Guide Semester Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Ccsd Chemistry 1 Study Guide Semester
 - Highlighting and Note-Taking Ccsd Chemistry 1 Study Guide Semester
 - Interactive Elements Ccsd Chemistry 1 Study Guide Semester
 8. Staying Engaged with Ccsd Chemistry 1 Study Guide Semester
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ccsd Chemistry 1 Study Guide Semester
 9. Balancing eBooks and Physical Books Ccsd Chemistry 1 Study Guide Semester
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ccsd Chemistry 1 Study Guide Semester
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Ccsd Chemistry 1 Study Guide Semester
 - Setting Reading Goals Ccsd Chemistry 1 Study Guide Semester
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Ccsd Chemistry 1 Study Guide Semester
 - Fact-Checking eBook Content of Ccsd Chemistry 1 Study Guide Semester
 - Distinguishing Credible Sources
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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
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 - Interactive and Gamified eBooks

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