

18.4

CALCULATIONS INVOLVING COLLIGATIVE PROPERTIES

SECTION REVIEW

Objectives

- Calculate the molality and mole fraction of a solution.
- Calculate the molal mass of a molecular compound from the freezing-point depression or boiling-point elevation of a solution of that compound.

Key Terms

- molality (m)
- mole fraction
- molal freezing-point depression constant (K_f)
- molal boiling-point elevation constant (K_b)

Key Equations

- Molality = $\frac{\text{moles of solute}}{\text{kilograms of solvent}} = \frac{\text{moles of solute}}{1000 \text{ g of solvent}}$
- mole fraction: $X_A = \frac{n_A}{n_A + n_B}$ $X_B = \frac{n_B}{n_A + n_B}$
where n_A = moles of solute
 n_B = moles of solvent
- $\Delta T_f = K_f m$
- $\Delta T_b = K_b m$

Part A Completion

After this completion exercise is done, you will demonstrate your understanding of the concepts and terms in this section by completing each block. Each block uses the completed block in terms of what phrase, or name, fits.

- Molality is an expression of concentration involving the ratio of _____ particles to _____ particles. Molality is expressed as moles of solute per _____ of solvent.
- An other expression of concentration is _____ in which concentrations are expressed as the ratio of moles of solute to the total number of moles of solute and solvent. Each solvent has a characteristic _____ elevation constant and molal freezing-point _____ constant. The elevation in boiling point of a solution can be calculated by multiplying the _____ concentration of the solute by the boiling-point _____ constant of the solvent.
- _____
 - _____
 - _____
 - _____
 - _____
 - _____
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 - _____

Calculations Involving Colligative Properties Section Review Answers

K Morrison



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