

7. When synthesizing soap, soap makers follow a formula to determine the weight of reagents to use. A value called the saponification number is used to determine the correct ratio of reagents. To produce a soap using coconut oil and sodium hydroxide, the saponification number is 191.1. Therefore, for every 1000 g of coconut oil used, 191.1 g of sodium hydroxide is required.

$$x \text{ grams NaOH} = x \text{ grams coconut oil} (191.1 \text{ grams NaOH} / 1000 \text{ grams coconut oil})$$

If you were to produce another batch of soap using 37.5 grams of coconut oil, how many grams of sodium hydroxide would you need? Show your calculations below.

Ex. $191.1 \text{ g NaOH} = 1000 \text{ g coconut oil} (191.1 \text{ g NaOH} / 1000 \text{ g coconut oil})$

$$\text{NaOH } 0.191 \text{ g} = 1000 \text{ g coconut oil} / 191.1 \text{ g NaOH}$$

$$\text{NaOH } 7.166 \text{ g} = 37.5 \text{ g coconut oil} (0.191 \text{ NaOH})$$

8. Using potassium hydroxide instead of sodium hydroxide yields softer soaps and creams. The saponification number of potassium hydroxide when combined with coconut oil is 268. Calculate how much potassium hydroxide one should add to 37.5 grams of coconut oil to produce soap.

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H. A. Appleton, W. H. Simmons



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