



The stronger acid is on the left. The equilibrium lies far to the right.

$$K_{eq} = \frac{10^{-34.9}}{10^{-36}} = 10^{1.1} = 1.2 \times 10^1$$

1.65 The conjugate base of squaric acid has the structure



Move electron pairs as indicated, beginning at the negatively charged oxygen.



1.79 By following the electron flow, express that the C—Br bond breaks with both electrons in that bond becoming an isolated pair of bromide ion. A new bond is formed between sulfur and the carbon from which the bromine departs. The electron pair responsible for this new bond is derived from the double bond between carbon and sulfur in the compound on the left (shown). The electron pair from the carbon to C—Br is compensated for by C=Br double bond formation resulting a total pair of nitrogen.



Answers to Interpretive Problems 4

1.71 B, 1.72 D, 1.73 A, 1.74 B, 1.75 A, 1.76 C

SELF-TEST

1. Write the electron configuration for each of the following:

(a) Phosphorus (b) Sulfide ion in Na_2S

2. Determine the formal charge of each atom and the net charge for each of the following species:

(a) $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$ (b) $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$ (c) $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$

3. Write a correct Lewis structure that satisfies the octet rule for each of the species in Problem

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