

Cracking the Periodic Table Code

Why aren't the elements listed in alphabetical order?

Why?

As charged, the periodic table is a bit odd. It's not square. Large portions of the table appear to be missing at the top. It's not organized alphabetically so elements can be found easily. But to a chemist, the periodic table is a very powerful tool. The periodic table is organized by properties, both chemical and physical. These properties relate to the electronic structure of the atoms of each element. In today's activities, you are going to study how the ground state electron configurations and structures of atoms are related to the shape and organization of the periodic table.

Model 1 – Blank Periodic Table

s-block														p-block														
1	1s ²														2s ²													
2	2s ²														2p ⁶													
3	3s ²														3p ⁶													
4	4s ²														4p ⁶													
5	5s ²														5p ⁶													
6	6s ²														6p ⁶													
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2	2p ⁶														3p ⁶													
3	3p ⁶														4p ⁶													
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6	6p ⁶														7p ⁶													
7	7p ⁶														8p ⁶													
8	8p ⁶														8p ⁶													

- Obtain a card with electron configurations for your group as assigned by your teacher.
- Record your name _____ and room number _____.
- For each element on your card, write the last orbital notation appearing at the end of the configuration (the underlined position).

Answers will vary for each atom. The answers for Team 3 are shown below.

Li 1s² Be 2s² B 2s² 2p¹
Li 2s² Be 2s² 2p¹ B 2s² 2p¹

- What is similar about the last orbital notation appearing at the end of the configuration for each element in your set?

Answers will vary for each atom. For Team 3, the configurations always end in s¹.

Cracking The Periodic Table Code Answers

Xiang Xie

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Atoms, Molecules & Elements: Patterns In the Periodic Table Gr. 5-8 George Graybill, 2015-10-01 This is the

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