

2.7 Ionic bonds are attractions between ions of opposite charge

In some cases, two atoms are so unequal in their attraction for electrons that the more electronegative atom strips an electron completely away from its partner. Table salt (NaCl) is an example of how this transfer of electrons can bond atoms together. Notice in **Figure 2.7A** that sodium has only 1 electron in its outer shell, whereas chlorine has 7. When these atoms interact, the sodium atom's outer electron is transferred to chlorine. Sodium now has only two shells, the second shell having a full set of 8 electrons. And chlorine's outer shell is also now full with 8 electrons.

But how does this electron transfer result in an ionic bond between Na and Cl ? Remember that electrons are negatively charged particles. The transfer of an electron moves one unit of negative charge from one atom to the other. Sodium, with 11 protons but now only 10 electrons, has a net electrical charge of $1+$. Chlorine, having gained an extra electron, now has 18 electrons but only 17 protons, giving it a net electrical charge of $1-$. In each case, an atom has become an **ion**—an atom or molecule with an electrical charge resulting from a gain or loss of one or more electrons. (Note that the names of negatively charged ions often end in *-ide*, such as chloride.) When the attraction between two ions with opposite charges holds them together, it is called an ionic bond. The resulting compound, in this case NaCl , is electrically neutral.

Sodium chloride is a familiar type of **salt**, a synonym for an ionic compound. Salts often exist as crystals in nature. **Figure 2.7B** shows the ions Na^+ and Cl^- in a crystal of sodium chloride. An NaCl crystal can be of any size

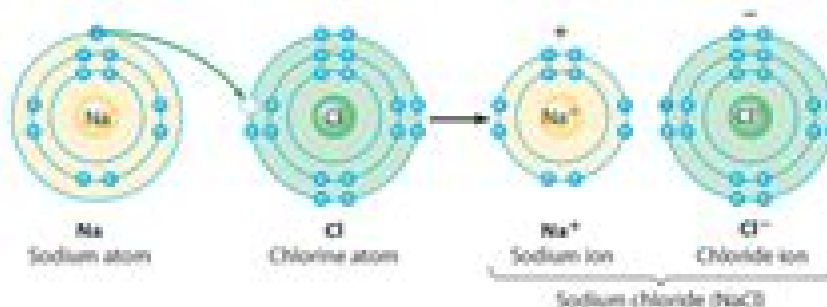


Figure 2.7A Electron transfer and the formation of an ionic bond

TRY THIS Explain why the sodium ion has a positive charge and the chloride ion has a negative charge.

(there is no fixed number of ions), but sodium and chloride ions are always present in a 1:1 ratio. The ratio of ions can differ in the various kinds of salts.

The environment affects the strength of ionic bonds. In a dry salt crystal, the bonds are so strong that it takes a hammer and chisel to break enough of them to crack the crystal. If the same salt crystal is placed in water, however, the ionic bonds break when the ions interact with water molecules and the salt dissolves, as we'll discuss in Module 2.13. Most drugs are manufactured as salts because they are quite stable when dry but can dissolve easily in water.



Figure 2.7B A crystal of sodium chloride

? Explain what holds together the ions in a crystal of table salt.

Write your answer in the space provided. You may use the space below to write your answer. Your answer will be graded by the system.



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the financial statement showing a firm s accounting value on a particular date is the