

# CHEMISTRY

## CHEAT SHEET

### CHEMICAL REACTIONS

#### CHEMICAL REACTIONS CHEAT SHEET

##### Analysis of a Chemical Reaction Equation



Type of Reaction	General Format
Combustion	$C_xH_y + O_2 \rightarrow CO_2 + H_2O$
Decomposition	$AB \rightarrow A + B$
Redox	One substance is oxidized (loses electrons) and another is reduced (gains electrons).
Acid-Base	$HA + B \rightarrow A^- + HB^+$
Double Displacement	$AB + CD \rightarrow AD + CB$
Redox	One substance is oxidized (loses electrons) and another is reduced (gains electrons).
Redox	$A + B \rightarrow A^+ + B^-$
Redox	$A + B \rightarrow A^+ + B^-$

##### Notation of a Chemical Reaction

- 1. Reactants
- 2. Products
- 3. Reaction arrow
- 4. Stoichiometric coefficients
- 5. State symbols
- 6. Reaction conditions

##### Stoichiometry and Chemical Reactions



##### Chemical Reaction Arrow Symbols

$\xrightarrow{\Delta}$	Reaction requires heat
$\xrightarrow{h\nu}$	Reaction requires light energy
$\xrightarrow{Pt}$	Reaction requires a platinum catalyst
$\xrightarrow{H^+}$	Reaction requires an acid catalyst
$\xrightarrow{OH^-}$	Reaction requires a hydroxide catalyst
$\xrightarrow{[O]}$	Oxidation reaction. The substance is oxidized. (Loss of electrons)
$\xrightarrow{[R]}$	Reduction reaction. The substance is reduced. (Gain of electrons)

##### Graphical Representation of Equilibrium and Kinetics



# Classifying Chemical Reactions Study Guide

**Arihant Experts**



## Classifying Chemical Reactions Study Guide:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. Organic Chemistry Reactions: A Study Guide Cybellium, 2024-10-26

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**Chemistry Study Guide** Robinson, 1992 Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon: buckyballs, graphite, and diamond are illustrated at the left, as is the molecule methane ( $\text{CH}_4$ ) from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement. *Basic Concepts of Chemistry, Study Guide*

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