

# Acids and Bases

Acid				Base			
Strength of Acid	Ka	Name	Formula	Formula	Name	Strength of Base	
<b>Strongest</b>						<b>Weakest</b>	
	Large	Perchloric acid	HClO <sub>4</sub>	ClO <sub>4</sub> <sup>-</sup>	Perchlorate ion		
	3.2×10 <sup>8</sup>	Hydroiodic acid		I <sup>-</sup>	Iodide		
	1.0×10 <sup>9</sup>	Hydrobromic acid	HBr	Br <sup>-</sup>	Bromide		
	1.3×10 <sup>9</sup>	Hydrochloric acid	HCl	Cl <sup>-</sup>	Chloride		
	1.0×10 <sup>3</sup>	Sulfuric acid	H <sub>2</sub> SO <sub>4</sub>	HSO <sub>4</sub> <sup>-</sup>	Hydrogen sulfate ion		
	2.4×10 <sup>1</sup>	Nitric acid	HNO <sub>3</sub>	NO <sub>3</sub> <sup>-</sup>	Nitrate ion		
	-----	Hydronium ion	H <sub>3</sub> O <sup>+</sup>	H <sub>2</sub> O	Water		
	5.4×10 <sup>-2</sup>	Oxalic acid	HO <sub>2</sub> C <sub>2</sub> O <sub>2</sub> H	HO <sub>2</sub> C <sub>2</sub> O <sub>2</sub> <sup>-</sup>	Hydrogen oxalate ion		
	1.3×10 <sup>-2</sup>	Sulfurous acid	H <sub>2</sub> SO <sub>3</sub>	HSO <sub>3</sub> <sup>-</sup>	Hydrogen sulfite ion		
	1.0×10 <sup>-2</sup>	Hydrogen sulfate ion	HSO <sub>4</sub> <sup>-</sup>	SO <sub>4</sub> <sup>2-</sup>	Sulfate ion		
	7.1×10 <sup>-3</sup>	Phosphoric acid	H <sub>3</sub> PO <sub>4</sub>	H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>	Dihydrogen phosphate ion		
	7.2×10 <sup>-4</sup>	Nitrous acid	HNO <sub>2</sub>	NO <sub>2</sub> <sup>-</sup>	Nitrite ion		
	6.6×10 <sup>-4</sup>	Hydrofluoric acid	HF	F <sup>-</sup>	Fluoride ion		
	1.8×10 <sup>-4</sup>	Methanoic acid	HCO <sub>2</sub> H	HCO <sub>2</sub> <sup>-</sup>	Methanoate ion		
	6.3×10 <sup>-5</sup>	Benzoic acid	C <sub>6</sub> H <sub>5</sub> COOH	C <sub>6</sub> H <sub>5</sub> COO <sup>-</sup>	Benzoate ion		
	5.4×10 <sup>-5</sup>	Hydrogen oxalate ion	HO <sub>2</sub> C <sub>2</sub> O <sub>2</sub> <sup>-</sup>	O <sub>2</sub> C <sub>2</sub> O <sub>2</sub> <sup>2-</sup>	Oxalate ion		
	1.8×10 <sup>-5</sup>	Ethanoic acid	CH <sub>3</sub> COOH	CH <sub>3</sub> COO <sup>-</sup>	Ethanoate (acetate) ion		
	4.4×10 <sup>-7</sup>	Carbonic acid	H <sub>2</sub> CO <sub>3</sub>	HCO <sub>3</sub> <sup>-</sup>	Hydrogen carbonate ion		
	1.1×10 <sup>-7</sup>	Hydrosulfuric acid	H <sub>2</sub> S	HS <sup>-</sup>	Hydrogen sulfide ion		
	6.3×10 <sup>-8</sup>	Dihydrogen phosphate ion	H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>	HPO <sub>4</sub> <sup>2-</sup>	Hydrogen phosphate ion		
	6.2×10 <sup>-8</sup>	Hydrogen sulfite ion	HS <sup>-</sup>	SO <sub>3</sub> <sup>2-</sup>	Sulfite ion		
	2.9×10 <sup>-8</sup>	Hypochlorous acid	HClO	ClO <sup>-</sup>	Hypochlorite ion		
	6.2×10 <sup>-10</sup>	Hydrocyanic acid	HCN	CN <sup>-</sup>	Cyanide ion		
	5.8×10 <sup>-10</sup>	Ammonium ion	NH <sub>4</sub> <sup>+</sup>	NH <sub>3</sub>	Ammonia		
	5.8×10 <sup>-10</sup>	Boric acid	H <sub>3</sub> BO <sub>3</sub>	H <sub>2</sub> BO <sub>3</sub> <sup>2-</sup>	Dihydrogen borate		
	4.7×10 <sup>-11</sup>	Hydrogen carbonate ion	HCO <sub>3</sub> <sup>-</sup>	CO <sub>3</sub> <sup>2-</sup>	Carbonate ion		
	4.2×10 <sup>-13</sup>	Hydrogen phosphate ion	HPO <sub>4</sub> <sup>2-</sup>	PO <sub>4</sub> <sup>3-</sup>	Phosphate ion		
	1.8×10 <sup>-13</sup>	Dihydrogen borate ion	H <sub>2</sub> BO <sub>3</sub> <sup>-</sup>	HBO <sub>3</sub> <sup>2-</sup>	Hydrogen borate ion		
	1.3×10 <sup>-13</sup>	Hydrogen sulfide ion	HS <sup>-</sup>	S <sup>2-</sup>	Sulfide ion		
	1.6×10 <sup>-14</sup>	Hydrogen borate ion	HBO <sub>3</sub> <sup>2-</sup>	BO <sub>3</sub> <sup>3-</sup>	Borate ion		
	-----	water	H <sub>2</sub> O	OH <sup>-</sup>	Hydroxide		
<b>Weakest</b>						<b>Strongest</b>	

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**Brian G. Cox**



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  - The Rise of Digital Reading **Download Table Of Acidbase Strength**
  - Advantages of eBooks Over Traditional Books
2. Identifying **Download Table Of Acidbase Strength**
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an **Download Table Of Acidbase Strength**
  - User-Friendly Interface
4. Exploring eBook Recommendations from **Download Table Of Acidbase Strength**
  - Personalized Recommendations
  - **Download Table Of Acidbase Strength** User Reviews and Ratings
  - **Download Table Of Acidbase Strength** and Bestseller Lists

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  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Download Table Of Acidbase Strength
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  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Download Table Of Acidbase Strength
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Download Table Of Acidbase Strength
  - Setting Reading Goals Download Table Of Acidbase Strength
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Download Table Of Acidbase Strength
  - Fact-Checking eBook Content of Download Table Of Acidbase Strength
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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