

20.12 Cyclohexane: a cyclic hydrocarbon

Physical properties and structure

Cyclohexane is a colourless, odourless liquid with a boiling point of 98.4°C and a melting point of 6.5°C. It is insoluble in water but miscible with most organic solvents. The molecule is non-polar and has a low dielectric constant. Cyclohexane is a saturated hydrocarbon, meaning it contains only single bonds between carbon atoms. The carbon atoms are arranged in a hexagonal ring, with two hydrogen atoms attached to each carbon atom, one above and one below the plane of the ring. This results in a chair conformation, which is the most stable arrangement of the atoms in the molecule.

Chemical properties and reactions

Cyclohexane is relatively inert to most chemical reactions. It does not react with strong acids or bases, and it is not easily oxidized. However, it can undergo free-radical halogenation in the presence of light or heat. For example, cyclohexane reacts with bromine to form bromocyclohexane. This reaction is a substitution reaction, where a hydrogen atom on the cyclohexane ring is replaced by a bromine atom.

Applications

Cyclohexane is used in a variety of industrial and laboratory applications. It is a common solvent for many organic compounds, particularly those that are non-polar. It is also used in the production of plastics and other polymers. In the laboratory, cyclohexane is used as a reagent in many chemical reactions, including the synthesis of cyclohexene and cyclohexanone.

Structure and properties of cyclohexane

Cyclohexane is a saturated hydrocarbon with the molecular formula C_6H_{12} . It consists of a hexagonal ring of six carbon atoms, each bonded to two hydrogen atoms.



Figure 20.12: Cyclohexane is a saturated hydrocarbon with the molecular formula C_6H_{12} . It consists of a hexagonal ring of six carbon atoms, each bonded to two hydrogen atoms.

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