

# Double Displacement Reactions

Directions for completing table below:

- Record your observations (Evidence of Reaction) of each experiment. Use the following terminology: (a) "Precipitate formed" (include the color), (b) "Gas evolved," (c) "Heat evolved," or (d) "No reaction observed."
- Complete and balance the equation for each case in which a reaction occurred. First write the correct formulas for the products, taking into account the charges (oxidation numbers) of the ions involved. Then balance the equation by placing a whole number in front of each formula (as needed) to adjust the number of atoms of each element so that they are the same on both sides of the equation. Use (g) or (s) to indicate gases and precipitates. Where no evidence of reaction was observed, write the words "No reaction" as the right-hand side of the equation.

Evidence of Reaction	Equation
1. no reaction	$\text{NaCl} + \text{KNO}_3 \longrightarrow \text{no reaction}$
2. precipitate	$\text{NaCl} + \text{AgNO}_3 \longrightarrow$
3. Evolution of gas	$\text{Na}_2\text{CO}_3 + \text{HCl} \longrightarrow$
4. Evolution of gas	$\text{NaOH} + \text{HCl} \longrightarrow$
5. precipitate	$\text{BaCl}_2 + \text{H}_2\text{SO}_4 \longrightarrow$
6. Evolution of gas	$\text{NH}_4\text{OH} + \text{H}_2\text{SO}_4 \longrightarrow$
7. no reaction	$\text{CuSO}_4 + \text{Zn}(\text{NO}_3)_2 \longrightarrow$
8. precipitate	$\text{Na}_2\text{CO}_3 + \text{CaCl}_2 \longrightarrow$
9. no reaction	$\text{CuSO}_4 + \text{NH}_4\text{Cl} \longrightarrow$
10. slight evolution of gas	$\text{NaOH} + \text{HNO}_3 \longrightarrow$
11. precipitate	$\text{FeCl}_3 + \text{NH}_4\text{OH} \longrightarrow$
12. Evolution of gas	$\text{Na}_2\text{SO}_3 + \text{HCl} \longrightarrow$

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