



Double and Triple Covalent Bonding – Examples



Covalent Bond: Bond between two non-metals.

Double Covalent Bond: Two non-metals bonding together by sharing 2 unpaired electrons per atom.

Triple Covalent Bond: Two non-metals bonding together by sharing 3 unpaired electrons per atom.

Atoms in the compound	Lewis Dot Diagram of the Final Compound	"Stick" Diagram
O_2 2 Oxygen Atoms	$\begin{array}{c} \cdot\ddot{O}\cdot \\ \cdot\ddot{O}\cdot \end{array} \rightarrow \begin{array}{c} \cdot\ddot{O}::\ddot{O}\cdot \\ \cdot\ddot{O}::\ddot{O}\cdot \end{array}$	$\begin{array}{c} \cdot\ddot{O}::\ddot{O}\cdot \\ \cdot\ddot{O}::\ddot{O}\cdot \end{array}$
N_2 2 Nitrogen Atoms	$\begin{array}{c} \cdot\ddot{N}\cdot \\ \cdot\ddot{N}\cdot \end{array} \rightarrow \begin{array}{c} \cdot\ddot{N}:::\ddot{N}\cdot \\ \cdot\ddot{N}:::\ddot{N}\cdot \end{array}$	$\begin{array}{c} \cdot\ddot{N}:::\ddot{N}\cdot \\ \cdot\ddot{N}:::\ddot{N}\cdot \end{array}$
CO_2 1 Carbon + 2 Oxygen	$\begin{array}{c} \cdot\ddot{C}\cdot \\ \cdot\ddot{O}\cdot \quad \cdot\ddot{O}\cdot \end{array} \rightarrow \begin{array}{c} \cdot\ddot{O}::\ddot{C}::\ddot{O}\cdot \\ \cdot\ddot{O}::\ddot{C}::\ddot{O}\cdot \end{array}$	$\begin{array}{c} \cdot\ddot{O}::\ddot{C}::\ddot{O}\cdot \\ \cdot\ddot{O}::\ddot{C}::\ddot{O}\cdot \end{array}$

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