

15. What is the diameter of a single iron atom in meters if it's diameter is $1.73 \times 10^{-5} \mu\text{m}$?

$$\frac{1.73 \times 10^{-5} \mu\text{m}}{1} \cdot \frac{1.0 \times 10^6 \text{ m}}{1 \mu\text{m}} = 1.73 \times 10^{-11} \text{ m}$$

16. How many joules does it take to raise the temperature of 3.50L of water from 22.0°C to 44.0°C ? $4,186 \text{ J/g}^\circ\text{C}$
How many kilojoules? Take the density of water to be 1.00g/mL .

$$\frac{1.00\text{g}}{\text{mL}} \cdot \frac{1000\text{mL}}{1\text{L}} = \frac{1000\text{g}}{\text{L}} \cdot 3.50\text{L} = 3500\text{g}$$

$$Q = \text{mass}(\text{water}) \cdot \text{specific heat of water} \cdot \Delta T$$

$$= 3500\text{g} \cdot \frac{4186\text{J}}{\text{g}^\circ\text{C}} \cdot (44.0 - 22.0)^\circ\text{C} = 321860\text{J} \cdot \frac{1\text{kJ}}{1000\text{J}} = 321.86\text{kJ}$$

17. Suppose 14.0g of magnesium (Mg) reacts with 63.0g of chloride to give 70.0g of the compound magnesium chloride (MgCl_2). In the process, all the magnesium gets used up, but some elemental chloride is left over.

a) For the law of conservation of mass to be obeyed, how much chloride is unused?

$$7.0\text{g}$$

b) What is the percent Mg in MgCl_2 ?

$$\frac{14.0\text{g Mg}}{70.0\text{g MgCl}_2} \times 100 = 20\%$$

c) What is the percent Cl in MgCl_2 ?

$$63.0\text{g} - 7.0\text{g} = 56.0\text{g Cl}$$

$$\frac{56.0\text{g Cl}}{70.0\text{g MgCl}_2} \times 100 = 80\%$$

d) What is the sum of the %Mg and %Cl in MgCl_2 ?

$$20\% + 80\% = 100\%$$

18. The compound carbon disulfide (CS_2) has a total mass of 76.15g. 12.01 g of carbon and 64.14g of sulfur. What is the mass of carbon in another sample of carbon disulfide, if the mass of sulfur is 21.38?

$$\frac{12.01\text{g C}}{76.15\text{g CS}_2} \times 100 = \frac{15.77\%}{100} \cdot 21.38\text{g CS}_2 = 3.372\text{g C}$$

19. A new element Uranium has 2 isotopes, Uranium 235 and Uranium 237. The isotopes masses are shown below.

Uranium-235 (234.99 amu, 78.09% abundance)

Uranium-237 (237.34 amu)

A. What is the percent abundance of the Uranium 237 isotope?

$$100\% - 78.09\% \text{ U-235} = 21.91\%$$

B. Calculate the weighted average of the atomic mass of naturally occurring uranium.

$$\left(234.99\text{amu} \cdot \frac{78.09\%}{100} \right) + \left(237.34\text{amu} \cdot \frac{21.91\%}{100} \right) =$$

$$183.50\text{amu} + 52.001 = 235.5\text{amu}$$

Calculating Percent Mass Volume Chem Worksheet 15 2

Answers

Albert Kaskel



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