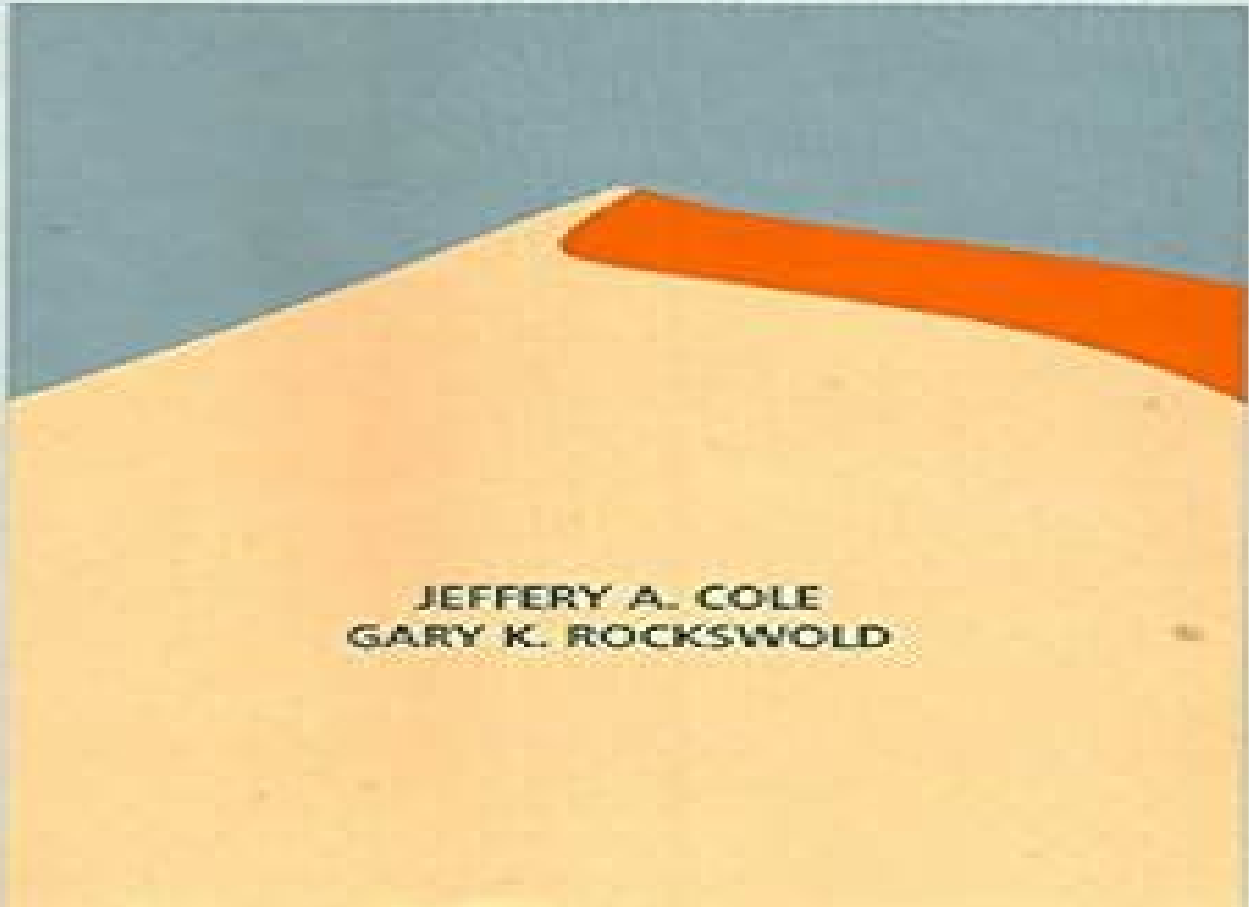


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VOL. 2**

SWOKOWSKI/OLINICK/PENCE
Calculus, Sixth Edition



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web this data shows that there was a gain the mass of the lighter which shows that there was a certain amount of butane 0 38 grams 0 02 grams gas that was successfully collected during the gas collection there is a mild influence of barometric pressure during both trials as shown above in the raw data

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of butane 0.23 g / 0.0040 moles = 58 g / mole. The molecular formula of butane is C_4H_{10} .

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