

Diagnostische toets klas 4 voorbereiding pw hoofdstuk 2

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Gebruik je boek en de Binas.

Schilkunde is ook leerwerk, in dit hoofdstuk worden heel wat begrippen genoemd en die moet je dus kennen.

1)

- a) Teken een Be-9 atoom volgens het model Rutherford/Bohr.
- b) De schillen worden met letters aangegeven, noteer de letters van de schillen die horen bij Be-9 en noteer het maximale aantal elektronen die deze schillen kunnen bevatten.
- c) Noteer hoeveel protonen, neutronen en elektronen een Be-9 atoom heeft.

2) Gegeven een atoomsoort met 18 elektronen.

Een leerling trekt de conclusie dat het om het atoomsoort Argon moet gaan.

Leg uit dat de leerling geen gelijk hoeft te hebben.

3)

- a) Leg uit waardoor de grootte van een atoom wordt bepaald.
- b) Leg uit waardoor de massa van een atoom wordt bepaald.

4) Schrijf een andere notatie voor: K-41

5) Wat geeft het atoomnummer in het periodiek systeem van een atoomsoort aan?

6) Het massagetal van een silicium-isotoop is 30.

- a) Leg uit waarom niet 30 u genoteerd mag worden.
- b) Op welke plaatsnummer in het periodiek systeem staat Si-30?
- c) Noteer het aantal valentie-elektronen van silicium.

7) Noteer de elektronenconfiguratie van yttrium.

8) Een atoomsoort staat in de vierde periode en behoort tot groep 16. Noteer de naam van dit atoomsoort.

9) Wat is de elektrovalentie oftewel waardigheid van gallium?

10) Hoeveel elektronen bevat een eenwaardig positief kaliumion.

11) Welke overeenkomst bestaat er (in het algemeen) tussen de atoomsoorten die onderelkaar staan in het periodiek systeem?

12) Noteer waar of niet-waar

- a) Isotopen van het atoomsoort Pb verschillen in chemische eigenschappen.
- b) Isotopen van het atoomsoort Pb verschillen in atoomnummer.
- c) Isotopen van het atoomsoort Pb verschillen in grootte.
- d) Isotopen van het atoomsoort Pb hebben gelijk aantal elektronen.
- e) Isotopen van het atoomsoort Pb verschillen in massa.
- f) Isotopen van het atoomsoort Pb verschillen in aantal protonen.
- g) Bij verwarmen van een stof worden de atomen een klein beetje groter.
- h) Tussen de elektronenbanen van een atoom zit lucht.
- i) Bij chemische reacties veranderen de elektronen.
- j) De atoomsoorten in groep 17 worden de halogenen genoemd.

Diagnostische Toets Hoofdstuk 11

A Loxley



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Table of Contents Diagnostische Toets Hoofdstuk 11

1. Understanding the eBook Diagnostische Toets Hoofdstuk 11
 - The Rise of Digital Reading Diagnostische Toets Hoofdstuk 11
 - Advantages of eBooks Over Traditional Books
2. Identifying Diagnostische Toets Hoofdstuk 11
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Diagnostische Toets Hoofdstuk 11
 - User-Friendly Interface
4. Exploring eBook Recommendations from Diagnostische Toets Hoofdstuk 11
 - Personalized Recommendations

- Diagnostische Toets Hoofdstuk 11 User Reviews and Ratings
- Diagnostische Toets Hoofdstuk 11 and Bestseller Lists
- 5. Accessing Diagnostische Toets Hoofdstuk 11 Free and Paid eBooks
 - Diagnostische Toets Hoofdstuk 11 Public Domain eBooks
 - Diagnostische Toets Hoofdstuk 11 eBook Subscription Services
 - Diagnostische Toets Hoofdstuk 11 Budget-Friendly Options
- 6. Navigating Diagnostische Toets Hoofdstuk 11 eBook Formats
 - ePub, PDF, MOBI, and More
 - Diagnostische Toets Hoofdstuk 11 Compatibility with Devices
 - Diagnostische Toets Hoofdstuk 11 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Diagnostische Toets Hoofdstuk 11
 - Highlighting and Note-Taking Diagnostische Toets Hoofdstuk 11
 - Interactive Elements Diagnostische Toets Hoofdstuk 11
- 8. Staying Engaged with Diagnostische Toets Hoofdstuk 11
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Diagnostische Toets Hoofdstuk 11
- 9. Balancing eBooks and Physical Books Diagnostische Toets Hoofdstuk 11
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Diagnostische Toets Hoofdstuk 11
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Diagnostische Toets Hoofdstuk 11
 - Setting Reading Goals Diagnostische Toets Hoofdstuk 11
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Diagnostische Toets Hoofdstuk 11
 - Fact-Checking eBook Content of Diagnostische Toets Hoofdstuk 11

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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