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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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 $\text{HCl (aq)} + \text{NaOH (aq)} \rightarrow \text{NaCl (aq)} + \text{H}_2\text{O (l)}$
 one mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4 \text{ aq} + 2\text{NaOH aq} \rightarrow \text{Na}_2\text{C}_2\text{O}_4 + 4\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a ph scale is used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction $\text{NaOH aq} + \text{C}_6\text{H}_4\text{COOH aq} \rightarrow \text{C}_6\text{H}_4\text{COONa aq} + \text{H}_2\text{O l}$ the point of equivalence is the point of titration when the number of moles of naoh na added is equal to the number of moles of khp nb in the