

1. Geef een complete uitwerking van de volgende redoxreacties:

A. Een H_2O_2 -oplossing wordt gevoegd bij een KI oplossing.

B. Zilver wordt opgelost in geconcentreerd salpeterzuur, hierbij ontstaat onder andere stikstofmonoxide.

C. Een natriumsulfiet-oplossing wordt toegevoegd aan een kaliumdichromaat-oplossing ($\text{K}_2\text{Cr}_2\text{O}_7$) waaraan een beetje verdund zwavelzuur is toegevoegd.

2. Men pipetteert 25,00 ml van een oxaalzuur ($\text{H}_2\text{C}_2\text{O}_4$)-oplossing, zuurt aan met zwavelzuur en titreert vervolgens met 0,504 M KMnO_4 -oplossing. Er blijkt 16,74 ml van deze oplossing nodig te zijn om de oplossing lichtroze te kleuren. Bepaal het massapercentage oxaalzuur. Ga ervan uit dat de dichtheid van de oxaalzuuroplossing $1,00 \text{ g.ml}^{-3}$ is.

3. Ammoniak in vervuilde lucht reageert met zwavelzuur tot ammoniumsulfaat. Via het regenwater komt deze stof in de bodem terecht. Onder invloed van bacteriën en in aanwezigheid van zuurstof en water worden de ammonium ionen omgezet in nitrietionen die op hun beurt weer verder reageren tot nitraationen.

A. Geef de vergelijking voor de vorming van ammoniumsulfaat in vervuilde lucht.

B. Is dit een redoxreactie? Leg uit.

C. Geef de vergelijking van de omzetting van ammoniumionen in nitrietionen (halfreactie)

D. Welke stof treedt op als oxidator? Geef de halfreactie.

E. Geef de vergelijkingen van de halfreacties en de somreactie van de omzetting van nitriet- in nitraationen.

F. Ammoniumnitriet ontleedt gemakkelijk in stikstof en water. Geef de reactievergelijking. Laat zien of dit een redoxreactie is.

Diagnostische Toets Redoxreacties

J Elliott



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Table of Contents Diagnostische Toets Redoxreacties

1. Understanding the eBook Diagnostische Toets Redoxreacties
 - The Rise of Digital Reading Diagnostische Toets Redoxreacties
 - Advantages of eBooks Over Traditional Books
2. Identifying Diagnostische Toets Redoxreacties
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Diagnostische Toets Redoxreacties
 - User-Friendly Interface
4. Exploring eBook Recommendations from Diagnostische Toets Redoxreacties
 - Personalized Recommendations

- Diagnostische Toets Redoxreacties User Reviews and Ratings
- Diagnostische Toets Redoxreacties and Bestseller Lists
- 5. Accessing Diagnostische Toets Redoxreacties Free and Paid eBooks
 - Diagnostische Toets Redoxreacties Public Domain eBooks
 - Diagnostische Toets Redoxreacties eBook Subscription Services
 - Diagnostische Toets Redoxreacties Budget-Friendly Options
- 6. Navigating Diagnostische Toets Redoxreacties eBook Formats
 - ePub, PDF, MOBI, and More
 - Diagnostische Toets Redoxreacties Compatibility with Devices
 - Diagnostische Toets Redoxreacties Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Diagnostische Toets Redoxreacties
 - Highlighting and Note-Taking Diagnostische Toets Redoxreacties
 - Interactive Elements Diagnostische Toets Redoxreacties
- 8. Staying Engaged with Diagnostische Toets Redoxreacties
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Diagnostische Toets Redoxreacties
- 9. Balancing eBooks and Physical Books Diagnostische Toets Redoxreacties
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Diagnostische Toets Redoxreacties
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Diagnostische Toets Redoxreacties
 - Setting Reading Goals Diagnostische Toets Redoxreacties
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Diagnostische Toets Redoxreacties
 - Fact-Checking eBook Content of Diagnostische Toets Redoxreacties

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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