

Combined Gas Law Practice 1.0 - Answer Key**Objective:** Define relationships between P, V, n, and T, and use them to solve gas law problems.**Directions:** Use $\left(\frac{P_1 V_1}{n_1 T_1}\right) = \left(\frac{P_2 V_2}{n_2 T_2}\right)$ to solve each gas law problem. Show your work!

1. The gas in a fixed container has a pressure of 1.20 atm and is at 245 K. What is the pressure of the gas if the temperature is increased to 299 K while moles and volume are constant?

$$\frac{P_1}{T_1} = \frac{P_2}{T_2} = \frac{1.20 \text{ atm}}{245 \text{ K}} = \frac{P_2}{299 \text{ K}}$$

$$\frac{1.20 \text{ atm} \times 299 \text{ K}}{245 \text{ K}} = P_2 = \boxed{1.51 \text{ atm}}$$

2. At constant temperature, 1.00 mol of a gas at 1.50 atm takes up 20.0 L. If the pressure changes to 2.50 atm, and a total of 2.00 moles, what has the volume changed to?

$$\frac{P_1 V_1}{n_1} = \frac{P_2 V_2}{n_2} = \frac{1.50 \text{ atm} \times 20.0 \text{ L}}{1.00 \text{ mol}} = \frac{2.50 \text{ atm} \times V_2}{2.00 \text{ mol}}$$

$$\frac{1.50 \text{ atm} \times 20.0 \text{ L} \times 2.00 \text{ mol}}{2.50 \text{ atm} \times 2.00 \text{ mol}} = V_2 = \boxed{12.0 \text{ L}}$$

3. A balloon with 0.55 moles of gas at 274 K takes up 12.3 L of space. If pressure and moles are constant, what will the volume of the balloon be when the temperature rises to 314 K?

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} = \frac{12.3 \text{ L}}{274 \text{ K}} = \frac{V_2}{314 \text{ K}}$$

$$\frac{12.3 \text{ L} \times 314 \text{ K}}{274 \text{ K}} = V_2 = \boxed{14.1 \text{ L}}$$

4. A fixed volume container has 1.20 moles of hydrogen gas at a pressure of 1.25 kPa and temperature of 325 K. What must the temperature be lowered to if we added in another mole of gas for a total of 2.20 moles, but need the pressure to fall to 0.95 kPa?

$$\frac{P_1}{n_1 T_1} = \frac{P_2}{n_2 T_2} = \frac{1.25 \text{ kPa}}{1.20 \text{ mol} \times 325 \text{ K}} = \frac{0.95 \text{ kPa}}{2.20 \text{ mol} \times T_2}$$

$$\frac{1.20 \text{ mol} \times 325 \text{ K} \times 0.95 \text{ kPa}}{0.95 \text{ kPa} \times 2.20 \text{ mol}} = T_2 = \boxed{199 \text{ K}}$$

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Sebastian Brünink

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