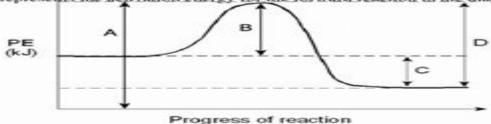
PART I Total Value: 50%

Instructions: Shade the letter of the correct answer on the computer scorable answer sheet provided.

Why does increasing temperature increase reaction rate?

(A) decreases activation energy

- (B) decreases the percentage of effective collisions
- (C) increases activation energy
- (D) increases the percentage of effective collisions
- 2. Which represents the activation energy for the forward reaction in the diagram below?



- (A) A (B) B
 - (C) C (D) D
- 3. Which type of change is best for monitoring the reaction rate of the reaction below?

$$Mg(OH)_2(s) + 2 HCl(aq) \rightarrow 2 H_2O(\ell) + MgCl_2(aq)$$

- (A) colour (B) mass
- (C) pH (D) volume
- 4. Which describes an activated complex?
 - (A) has a lower potential energy than products (B) has a lower potential energy than reactants
- (C) stable species that forms upon collision of reactants
 (D) unstable species that forms upon collision of reactants
- 5. According to the collision theory, why does the reaction below occur in more than one step?

$$C_5H_{12}(g) + 8 O_2(g) \rightarrow 5 CO_2(g) + 6 H_2O(g)$$

- (A) low C₃H₁₂(g) concentration
 - (B) low probability of a nine-particle collision
 - (C) particles collide with insufficient kinetic energy
 - (D) temperature is very high

Collision Theory Chemistry Answer Key

AW Rasmussen

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- 1. Understanding the eBook Collision Theory Chemistry Answer Key
 - The Rise of Digital Reading Collision Theory Chemistry Answer Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Collision Theory Chemistry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Collision Theory Chemistry Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Collision Theory Chemistry Answer Key
 - Personalized Recommendations

- Collision Theory Chemistry Answer Key User Reviews and Ratings
- Collision Theory Chemistry Answer Key and Bestseller Lists
- 5. Accessing Collision Theory Chemistry Answer Key Free and Paid eBooks
 - Collision Theory Chemistry Answer Key Public Domain eBooks
 - Collision Theory Chemistry Answer Key eBook Subscription Services
 - Collision Theory Chemistry Answer Key Budget-Friendly Options
- 6. Navigating Collision Theory Chemistry Answer Key eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Collision Theory Chemistry Answer Key Compatibility with Devices
 - Collision Theory Chemistry Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Collision Theory Chemistry Answer Key
 - Highlighting and Note-Taking Collision Theory Chemistry Answer Key
 - Interactive Elements Collision Theory Chemistry Answer Key
- 8. Staying Engaged with Collision Theory Chemistry Answer Key
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Collision Theory Chemistry Answer Key
- 9. Balancing eBooks and Physical Books Collision Theory Chemistry Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Collision Theory Chemistry Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Collision Theory Chemistry Answer Key
 - Setting Reading Goals Collision Theory Chemistry Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Collision Theory Chemistry Answer Key
 - Fact-Checking eBook Content of Collision Theory Chemistry Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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