

Lab 00 Report- Exploring the Properties of Gasses

CHEM 1225

Introduction

Throughout scientific history, a number of laws have been proposed to explain the properties of gasses under specific conditions. These gas laws incorporate the variables of pressure, volume, and temperature to predict the behavior and reactivity of a gaseous substance. Charles' Law states that the volume of a set quantity of gas is directly proportional to temperature. Under constant temperature, Boyle's Law suggests a direct relationship between volume and pressure. Gay-Lussac's Law relates temperature to pressure in a gas, and finally, Avogadro's Law explains the relationship between volume and the number of molecules (under constant temperature and pressure) in a gaseous substance.

These three laws combine to form the Ideal Gas Law ($PV = nRT$), which relates pressure, temperature, volume, number of molecules, and a figure called the *ideal gas constant* (approximately $8.315 \text{ J} \cdot \text{mol}^{-1} \cdot \text{K}^{-1}$). It is possible to determine an unknown value pertaining to a gas by identifying your known variables, "plugging" them into the Ideal Gas Law, and solving for the unknown variables.

Purpose

The main objective of this lab was to conduct a series of experiments that demonstrate each gas law. By varying the conditions of the gas in terms of pressure, temperature, and volume, it is possible to identify trends that align with known gas laws.

Procedure

Part I of the experiment started with a plastic, 20 mL syringe whose piston was positioned so that a full 20 mL of air was trapped in the barrel. The syringe was then attached to a gas pressure sensor, which was also attached to a LabQuest for real-time data collection. The LabQuest was set up to read volume (in mL) and pressure (in kPa). Pressure was adjusted by hand using the syringe. Data collection was initiated, the syringe was positioned, and the pressure was recorded as the volume of the syringe was decreased by an interval of 2 mL at each data collection point. After reaching 8 mL of remaining volume in the syringe, data collection was halted and the results were recorded.

To start Part II of the experiment, a Erlenmeyer flask with a rubber stopper was submerged in a water bath with variable temperature. A gas pressure sensor and temperature probe were connected to a LabQuest, which was used to measure the temperature of the water bath and the pressure within the flask. Beginning with ice water, the pressure and temperature was recorded as evidenced by the LabQuest. As the water in the bath naturally warmed to room temperature, pressure and temperature were recorded at intervals of 2 degrees C. After six total entries, data collection was halted and results were recorded.

Exploring The Properties Of Gases Lab Report

Arturo Cuomo



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