Name	

Date \_\_\_\_\_

# **Pre-laboratory Assignment**

- 1. How does temperature affect the kinetic energy of gas molecules?
- 2. What are the names and equations for the three gas laws tested in this lab?

- 3. Complete the following statements:
  - a. When T is constant and volume increases, the pressure \_\_\_\_\_
  - b. When T is constant and volume decreases, the pressure \_\_\_\_
  - c. When V is constant and temperature increases, the pressure \_\_\_\_
  - d. When V is constant and temperature decreases, the pressure
  - e. When P is constant and temperature increases, the volume \_
  - f. When P is constant and temperature decreases, the volume
  - 4. A sample of helium has a volume of 325 mL and a pressure of 655 mm Hg. What will be the pressure if the helium is compressed to 125 mL (T constant)?

5. A gas with a volume of 525 mL at a temperature of -25°C is heated to 75°C. What is the new volume of the gas if pressure and quantity of gas are held constant?

# **Exploring The Properties Of Gases Lab**

**Oak Ridge National Laboratory** 

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