

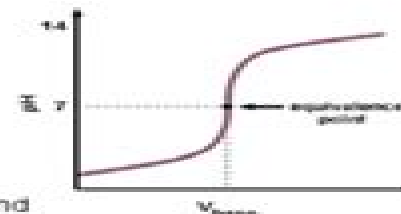
GHS Honors Chemistry Acid-Base Titration Lab

Name: _____ Date: _____ Block: _____

PURPOSE: To determine the concentration of a solution of hydrochloric acid by acid-base titration.

INTRODUCTION:

Titration is a laboratory technique by which we can determine the concentration of an unknown solution using a standard (known) concentration of another solution that chemically reacts with the unknown. This standard solution is called the **titrant**. We will also need a way to know that the reaction between the unknown solution, and the standard solution, is complete. In acid/base titrations, an indicator, much like those that you have already seen in the lab, are used. When the indicator changes color (colorless to pink, or vice versa), this indicates that the **equivalence point** is reached. At the end point:



$$N_{\text{titrant}} \times V_{\text{titrant}} = N_{\text{unknown}} \times V_{\text{unknown}}$$

In an Acid/Base titration, the Normality (N) is simply:

$$N = M \times n$$

where M = the Molarity of the solution, and n = times the number of or equivalents (the # of H⁺ or OH⁻ ions that dissociate from the acid or base).

For example:

- HCl contains only one H⁺, so n = 1, and the Normality (N) = the Molarity (M)
- H₂SO₄ contains two H⁺, so n = 2, and the Normality (N) = 2 x M
- NaOH contains only one OH⁻, so n = 1, and the Normality (N) = 1 x M
- Ca(OH)₂ contains two OH⁻, so n = 2, and the Normality (N) = 2 x M

The calculation for an Acid/Base Titration can be reduced to the following:

$$n_{\text{acid}} \times M_{\text{acid}} \times V_{\text{acid}} = n_{\text{base}} \times M_{\text{base}} \times V_{\text{base}}$$

In this experiment you will titrate a measured volume of HCl with a solution of NaOH of known concentration. The acid and the base react with one another according to the equation:



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