CH 15 Equilibrium Study Guide

Dynamic Equilibrium

Dynamic Equilibrium - forward & reverse reaction rates are equal

- ASSUME all rxns are ELEMENTAL (1 step!)
- Reversible reaction something constantly produced & consumed
 - Indicated by

No net change in reactant or product concentrations

Equilibrium Constant, Ke

For a general equation:

$$aA + bB \xrightarrow{k_t} aD + eE$$

$$K_c = \frac{k_t}{k_c} = \frac{[D]^d [E]^e}{[A]^e [B]^b}$$

A. The equilibrium-constant depends only on the stoichiometry of the reaction, not the mechanism.

B. DOES NOT depend on initial concentrations because only the RATIO of concentrations at equilibrium is the same

C. K has NO units!

D. ONLY consider gas & aqueous phases!!

-OMIT solids & liquids

Significance of K

Value of K indicates if more products or reactants are present at equilibrium

- . If K is large (>1): Products favored Equilibrium lies to the right
- . If K is small (<1): Reactants favored Equilibrium lies to the left

Equilibrium Constant, Ko

For a general equation:

Only if ALL phases are gases!

$$aA + bB \xrightarrow{k_f} dD + eE K_p = \frac{(P_D)^r(P_E)^r}{(P_A)^r(P_B)^r}$$

Relating Ko and Ko

$$K_{p} = K_{c}(RT)^{\Delta n}$$
; $\Delta n = [\Sigma \text{ (moles GAS in product)} - \Sigma \text{ (moles GAS in reactant)}]$

Dynamic Equilibrium Study Guide

Lauren Gardner

Dynamic Equilibrium Study Guide:

Reviewing **Dynamic Equilibrium Study Guide**: Unlocking the Spellbinding Force of Linguistics

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