Answers to Sample Questions for the Chemistry Placement Exam

Note: These questions are intended to represent the general level of difficulty and areas of coverage of the Chemistry Placement Exam.

They are NOT intended to be a complete representation of the specific types of questions to be found on the exam.

1.	Aluminum reacts with element X to form a compound with the formula Al ₂ X ₃ . Elemen X exists as diatomic molecules, X ₂ , in the gaseous state at normal temperature and pressure. Element X must be
a)	Nitrogen
Answer b)	Oxygen
c)	Sulfur
d)	Chlorine
Explanation:	Aluminum forms a +3 ion. In order for the formula unit Al_2X_3 to be electrically neutral, the element X must form a -2 ion. Nonmetals in the oxygen column do this. Of these elements, only oxygen is a diatomic gas.
2.	What is the coefficient of O_2 when the following equation is balanced? $C_2H_4O + O_2 \rightarrow CO_2 + H_2O$
a)	2
b)	3
c)	4
Answer d)	5
Explanation:	The balanced equation is 2 C₂H₄O + 5 O₂ → 4 CO₂ + 4 H₂O

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