

Formula of an Ionic Compound

Balancing Charges on Ions

Introduction

Atoms of different elements combine with one another to form compounds. The empirical formula of an ionic compound indicates the kinds of atoms that are present in the compound, as well as the relative number (ratio) of each kind of atom. Let's investigate how the formula of an ionic compound can be determined experimentally.

Concepts

- Ionic compounds
- Empirical formula
- Polyatomic ions
- Precipitation reaction

Background

An ionic compound is composed of ions—atoms or groups of atoms that have a positive or negative charge. Oppositely charged ions arrange themselves into an extended, three-dimensional structure called a crystal lattice. The net attractive forces among oppositely charged ions in the crystal structure are called ionic bonds. Although composed of charged ions, ionic compounds are electrically neutral. The ratio of oppositely charged ions in the crystal structure is such that the positive charge contributed by the cations is equal to or balanced by the negative charge contributed by the anions. There is no net or overall charge in an ionic compound.

The *empirical formula* of an ionic compound indicates the smallest whole-number ratio of each type of ion in the crystal structure and is called a *formula unit*. For example, magnesium chloride has the empirical formula MgCl_2 . Magnesium cations (Mg^{2+}) and chloride anions (Cl^-) combine in a 1:2 ratio to form the MgCl_2 formula unit. The overall charge on ionic compounds is always zero.

Some ions consist of a charged group of covalently bonded atoms. Such ions are called *polyatomic ions*. An example is the nitrate ion (NO_3^-), which contains one nitrogen atom and three oxygen atoms, and has an overall charge of -1 . In calcium nitrate, calcium (Ca^{2+}) ions combine with nitrate ions in a 1:2 ratio in order to balance the positive and negative charges. The empirical formula for calcium nitrate is $\text{Ca}(\text{NO}_3)_2$. Parentheses are used around the nitrate ion to show that the subscript, “2” pertains to the nitrate ion as a whole.

Many ionic compounds can be prepared in the lab using *precipitation reactions*. When solutions of two ionic compounds are combined, the ions may rearrange to form a new ionic compound that is insoluble in water. An example of this type of reaction is the formation of solid barium sulfate when barium chloride and sodium sulfate are combined in solution (Equation 1a). In Equation 1b, only the ions that form the precipitate are represented. This makes it easier to recognize what happens in the precipitation reaction.



According to the balanced equation for this reaction, barium ions (Ba^{2+}) combine with sulfate ions (SO_4^{2-}) in a 1:1 ratio to form barium sulfate (BaSO_4). This ratio can be observed experimentally in the lab by mixing $\text{BaCl}_2(\text{aq})$ and $\text{Na}_2\text{SO}_4(\text{aq})$ solutions containing equal amounts (concentration) of barium and sulfate ions, respectively. The maximum amount of precipitate will be obtained when equal volumes (a 1:1 ratio) of the two solutions are combined. A similar approach can also be used to determine the formula of an unknown ionic compound.

Experiment Overview

The purpose of this experiment is to determine the empirical formula of an unknown ionic compound. Two solutions containing equal amounts (concentration) of two reacting ions will be combined in a series of reactions. In each reaction, the total volume of the two solutions will be held constant while the volume ratio of the reactants is varied. The amount of precipitate obtained in each reaction will be measured and plotted against the volume ratio to find the empirical formula of the product.

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Crispin Collins



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