

Lesson 2: Stoichiometry

8. Show that you identify an ACID by looking at its formula?

Acids BEGINS with "H"

9. Show that you identify a BASE by looking at its formula, except for NH_3 (ammonia)?

Bases ENDS with "OH"

10. List the 8 strong acids.

Formula	HA	HCl	HBr	HNO_3	H_2SO_4	HClO_4
Name	Hydrofluoric acid	Hydrochloric acid	Hydrobromic acid	nitric acid	sulfuric acid	perchloric acid

11. Complete the sentence: Strong acids contain a cation from an element in group # _____, or _____, AND an anion called _____ with the symbol _____.

12. What is another term for "precipitation" reactions? Double displacement (met) reactions

13. Define precipitation: a solid that forms during a reaction and doesn't dissolve in water

14. Aqueous solutions contain ions that are present in solutions before and after a reaction. What are they called?

15. Predict the products of each reaction. Write a net ionic equation. Tell whether a reaction is likely to occur. (Use a solubility chart when necessary). Finally, balance the net ionic equations.

a) $\text{HCl(aq)} + \text{Zn(OH)}_2\text{(aq)} \rightarrow \text{ZnCl}_2\text{(aq)} + \text{Zn(OH)}_2$ Because all chemicals are aqueous and there are no weak electrolytes, this reaction will have NO NET ionic equation. (NO ions will separate and not be equalized in the equation). This means NO REACTION occurs.

b) $\text{CaCl}_2\text{(aq)} + \text{Na}_2\text{CO}_3\text{(aq)} \rightarrow \text{Na}_2\text{Cl}_2\text{(aq)} + \text{CaCO}_3\text{(s)}$ The Ionic Equation: $\text{Ca}^{2+} + \text{CO}_3^{2-} \rightarrow \text{CaCO}_3$ This reaction is likely to occur because there is a net ionic equation.

c) $\text{H}_3\text{PO}_4\text{(aq)} + \text{AlCl}_3\text{(aq)} \rightarrow \text{AlPO}_4\text{(s)} + \text{HCl(aq)}$ The Ionic Equation: $\text{Al}^{3+} + \text{PO}_4^{3-} \rightarrow \text{AlPO}_4$ This reaction is likely to occur because there is a net ionic equation.

16. Balance the equations for the following precipitation reactions, and then write the net ionic equations. Indicate the state of the species (s, l, aq, or g).



17. Oxalic acid, $\text{H}_2\text{C}_2\text{O}_4$, which is found in various plants, can provide two hydrogen ions to water. Write one balanced equation showing the acid losing its first hydrogen ion. Then, write a second equation showing the behavior of the oxalate acid losing its second hydrogen ion.



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