

Chapter 10: Gas Laws (Book problem numbers may)

Date _____

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Computation involving calculations and the given information involving the formulae at bottom for the unknown and plugging in the information/number and solving and writing the answer algebraically.

1. A sample of 0.800 moles of gas is collected in a container at volume of 0.010 L. What is the pressure of the gas at 200 K if the gas is at 20 °C? (book problem 10.1)

$$P = \frac{nRT}{V} = \frac{(0.800 \text{ moles})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(200 \text{ K})}{0.010 \text{ L}} = 13136 \text{ atm}$$

2. If 0.010 moles of gas are collected at 200 K, at 1000 atm, what volume of gas will be collected at 20 °C?

$$V = \frac{nRT}{P} = \frac{(0.010 \text{ moles})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(200 \text{ K})}{1000 \text{ atm}} = 0.00164 \text{ L}$$

3. Calculate the density of a gas at 1000 atm and 200 K. (book problem 10.2)

$$\rho = \frac{PM}{RT} = \frac{(1000 \text{ atm})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})}{(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(200 \text{ K})} = 50.0 \text{ g/L}$$

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4. If 0.010 moles of gas are collected at 1000 atm, at 200 K, what volume of gas will be collected at 20 °C? (book problem 10.3)

$$V = \frac{nRT}{P} = \frac{(0.010 \text{ moles})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(200 \text{ K})}{1000 \text{ atm}} = 0.00164 \text{ L}$$

5. Calculate the volume of 0.010 moles of gas at 1000 atm, at 200 K. (book problem 10.4)

$$V = \frac{nRT}{P} = \frac{(0.010 \text{ moles})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(200 \text{ K})}{1000 \text{ atm}} = 0.00164 \text{ L}$$

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$$V = \frac{nRT}{P} = \frac{(0.010 \text{ moles})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(200 \text{ K})}{1000 \text{ atm}} = 0.00164 \text{ L}$$

6. If 0.010 moles of gas are collected at 1000 atm, at 200 K, what volume of gas will be collected at 20 °C? (book problem 10.5)

$$V = \frac{nRT}{P} = \frac{(0.010 \text{ moles})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(200 \text{ K})}{1000 \text{ atm}} = 0.00164 \text{ L}$$

Ch 11 Gases Review Answers

Brigitte Niedzwiecki

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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