

## Physical Science – Chapter 18 Study Guide Answers

### Chapter Objectives

1. Define pure substances and mixtures.

**Pure substances – type of matter with a fixed composition, element or compound (composition is definite)**

**Mixtures – 2 or more substances that can be separated by physical means (Composition is variable)**

2. Identify elements and compounds.

**Elements – pure substances in which all atoms have the same identity**

**Compounds – when 2 or more different elements combine in a fixed proportion**

3. Compare and contrast solutions, colloids, and suspensions.

**Solutions – particles are so small they cannot be seen with a microscope and will never settle out; light cannot pass through, homogenous mixtures**

**Colloids – particles that are larger than those in solution but not heavy enough to settle out; Ex: milk, light passes through/scatters (Tyndall effect)**

**Suspensions – heterogeneous mixtures, contains liquid, particles will settle if left undisturbed, Ex: pond water; will scatter light**

4. Describe the phase and energy changes associated with STATE changes.

**Changing states of matter (from solids, liquids, or gas to another state), physical changes**

**Melting – solid to a liquid, add energy**

**Freezing: liquid to solid, releases energy**

**Vaporization: liquid to gas, add energy**

- Boiling: throughout the liquid

- Evaporation: surface only

**Condensation: gas to liquid, releases energy**

**Sublimation: solid to gas, adds energy**

5. Identify substances using physical properties.

**Physical properties – properties than can be observed without altering identity of substance**

**Alterations in form but not identity**

**Density, freezing/melting point, color**

6. Compare and contrast physical and chemical changes.

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**SA Dillow**



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