

4. a.
$$\frac{10g \text{ Al} \mid 1 \text{ mol Al} \mid 3 \text{ mol Cu} \mid 63.55g \text{ Cu}}{26.98g \text{ Al} \mid 2 \text{ mol Al} \mid 1 \text{ mol Cu}} = 35.33g \text{ Cu}$$

$$\frac{20g \text{ CuCl}_2 \mid 1 \text{ mol CuCl}_2 \mid 3 \text{ mol Cu} \mid 63.55g \text{ Cu}}{134.45g \text{ CuCl}_2 \mid 3 \text{ mol CuCl}_2 \mid 1 \text{ mol Cu}} = \boxed{9.45g \text{ Cu}}$$

b.
$$\frac{30g \text{ Al} \mid 1 \text{ mol Al} \mid 2 \text{ mol AlCl}_3 \mid 133.33g \text{ AlCl}_3}{26.98g \text{ Al} \mid 2 \text{ mol Al} \mid 1 \text{ mol AlCl}_3} = 148.25g \text{ AlCl}_3$$

$$\frac{100g \text{ CuCl}_2 \mid 1 \text{ mol CuCl}_2 \mid 2 \text{ mol AlCl}_3 \mid 133.33g \text{ AlCl}_3}{134.45g \text{ CuCl}_2 \mid 3 \text{ mol CuCl}_2 \mid 1 \text{ mol AlCl}_3} = \boxed{66.11g \text{ AlCl}_3}$$

c.
$$\frac{5g \text{ Al} \mid 1 \text{ mol Al} \mid 3 \text{ mol Cu} \mid 63.55g \text{ Cu}}{26.98g \text{ Al} \mid 2 \text{ mol Al} \mid 1 \text{ mol Cu}} = 17.67g \text{ Cu}$$

~~10g / 17.67g~~ $10g / 17.67g \times 100 = \boxed{56.6\%}$

d.
$$\frac{10g \text{ CuCl}_2 \mid 1 \text{ mol CuCl}_2 \mid 2 \text{ mol AlCl}_3 \mid 133.33g \text{ AlCl}_3}{134.45g \text{ CuCl}_2 \mid 3 \text{ mol CuCl}_2 \mid 1 \text{ mol AlCl}_3} = 6.61g \text{ AlCl}_3$$

~~5g / 6.61g~~ $5g / 6.61g \times 100 = \boxed{75.6\%}$

5. a. TRUE
b. FALSE
c. $\text{AgNO}_3 \Rightarrow 169.91g/\text{mol}$
d. 6.02×10^{23} particles = 1 mole of particles

Ch 11 Stoichiometry Study Guide

N Colangelo



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Table of Contents Ch 11 Stoichiometry Study Guide

1. Understanding the eBook Ch 11 Stoichiometry Study Guide
 - The Rise of Digital Reading Ch 11 Stoichiometry Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Ch 11 Stoichiometry Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ch 11 Stoichiometry Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ch 11 Stoichiometry Study Guide
 - Personalized Recommendations
 - Ch 11 Stoichiometry Study Guide User Reviews and Ratings
 - Ch 11 Stoichiometry Study Guide and Bestseller Lists
5. Accessing Ch 11 Stoichiometry Study Guide Free and Paid eBooks
 - Ch 11 Stoichiometry Study Guide Public Domain eBooks
 - Ch 11 Stoichiometry Study Guide eBook Subscription Services

- Ch 11 Stoichiometry Study Guide Budget-Friendly Options
- 6. Navigating Ch 11 Stoichiometry Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Ch 11 Stoichiometry Study Guide Compatibility with Devices
 - Ch 11 Stoichiometry Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ch 11 Stoichiometry Study Guide
 - Highlighting and Note-Taking Ch 11 Stoichiometry Study Guide
 - Interactive Elements Ch 11 Stoichiometry Study Guide
- 8. Staying Engaged with Ch 11 Stoichiometry Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ch 11 Stoichiometry Study Guide
- 9. Balancing eBooks and Physical Books Ch 11 Stoichiometry Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ch 11 Stoichiometry Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ch 11 Stoichiometry Study Guide
 - Setting Reading Goals Ch 11 Stoichiometry Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ch 11 Stoichiometry Study Guide
 - Fact-Checking eBook Content of Ch 11 Stoichiometry Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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