

EXPERIMENT 19 Calcium Compounds and Le Chatelier's Principle

This chemical reaction had not yet reached equilibrium, so the color changed over time. This reaction is summarized below:



Procedure:

Watch the two tubes versus earlier about Calcium Compound tests, equilibrium. Complete this table below, identify, present the equilibrium shift below:

https://www.youtube.com/watch?v=K256K23_78

Data and Discussion:

19. EFFECT OF CONCENTRATION		
Substances added	Observations	Shift in equilibrium (to the right, to the left, no shift)
+ Concentrated HCl	When this concentrated HCl was added, the pink solution became blue almost instantly. This produced a solution slightly cloudy on the sides, appearing as if the blue color was fading. The solution layer at the bottom faded, and the color faded, while the upper layer remained pink.	Shift to right
+ CaCl_2	When this concentrated CaCl_2 was added, the solution became pink. The color faded, and the color faded, while the upper layer remained pink.	Shift to right
+ $\text{H}_2\text{C}_2\text{O}_4$	When this concentrated $\text{H}_2\text{C}_2\text{O}_4$ was added, the solution became pink. The color faded, and the color faded, while the upper layer remained pink.	Shift to left
	When this HCl was added, the solution became pink. The color faded, and the color faded, while the upper layer remained pink.	Shift to right

Experiment 9 Chemical Equilibrium

Richard Bailey



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