

Experiment 8 Report Sheet

Limiting Reactant

Date _____

A. Precipitation of $\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$ from the Salt Mixture

Unknown number B

1. Mass of beaker (g)
2. Mass of beaker and salt mixture (g)
3. Mass of salt mixture (g)
4. Mass of filter paper (g)
5. Mass of filter paper and product after air-dried or oven-dried (g)
6. Mass of dried product (g) 1.191
7. Formula of dried product

B. Determination of Limiting Reactant

1. Limiting reactant in salt mixture (write complete formula)
2. Excess reactant in salt mixture (write complete formula)

Data Analysis

1. Moles of $\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$ (or CaC_2O_4) precipitated (mol)
2. Moles of limiting reactant in salt mixture (mol)
• formula of limiting hydrate $\text{K}_2\text{C}_2\text{O}_4 \cdot \text{H}_2\text{O}$
3. Mass of limiting reactant in salt mixture (g)
• formula of excess hydrate $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$
4. Mass of excess reactant in salt mixture (g)
• formula of excess hydrate $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$
5. Percent limiting reactant in salt mixture (%)
6. Percent excess reactant in salt mixture (%)
7. Mass of excess reactant that reacted (g)
8. Mass of excess reactant, unreacted (g)

*Show calculations for Trial 1 on next page.

Trial 1 250 mL

Trial 2

96.553g

47.956g

97.559g

49.083g

1.001g

1.067g

0.785g

0.785g

1.976g

1.774g

1.191g

0.989g

$\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$

$\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$

$\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$

$\text{K}_2\text{C}_2\text{O}_4 \cdot \text{H}_2\text{O}$

0.00815 mol

0.00676 mol

2. Moles of limiting reactant in salt mixture (mol)

• formula of limiting hydrate $\text{K}_2\text{C}_2\text{O}_4 \cdot \text{H}_2\text{O}$

0.00815 mol

0.00676 mol

3. Mass of limiting reactant in salt mixture (g)

1.198g CaCl_2

0.994g CaCl_2

4. Mass of excess reactant in salt mixture (g)

0.192g

0.073g

5. Percent limiting reactant in salt mixture (%)

93.97%

93.16%

6. Percent excess reactant in salt mixture (%)

19.32%

6.84%

7. Mass of excess reactant that reacted (g)

1.4922g

1.2343g

8. Mass of excess reactant, unreacted (g)

1.3030g

1.1606g

$$\textcircled{D} \quad \frac{1.191g \text{ CaC}_2\text{O}_4}{146.12g} \times 1 \text{ mol}$$

$$\frac{0.989g}{146.12g} \times 1 \text{ mol}$$

Experiment 8 Limiting Reactant Answers

Mary Millhollon, Richard H. Langley



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